# Unit #1

# Chapter #13: Physical Properties of Solutions

- What are the Types of Solutions?
  - Solution homogeneous mixture



- The nubbing alcohol sold in drug stores often is composed of 70% isopropyl alcohol and 30% water.
   In this solution
  - 60 water is the solvent.
  - (80) sopropyll alcohol is the solvent.
  - to both water and isopropyl alcohol are solvents.
  - D) neither water nor isopropyl alcohol is a solvent.



Unsaturated, Saturated, Supersaturated?

Unsaturated → less than dissolving capacity
Saturated → reaching max dissolving amount

- Supersaturated -> More is dissolved than allowed (high temp cooled)
- 11) An unsaturated solution is one that \_
  - A) has a concentration lower than the solubility
  - B) has no double bonds
  - C) contains the maximum concentration of solute possible, and is in equilibrium with undissolved solute
     D) contains no solute
- E) contains more dissolved solute than the solubility allows
  - What are the Solution Units?

Molarity (M) = moles of solution (L)

Molality (m) = moles of solution

Molality (m) = moles of solve (KB)

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**N** Noddings

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