

Gas Law Formula

Gas Law	Formula	Description
Boyle's Law	$P_1V_1 = P_2V_2$	At constant T , as pressure increases, volume decreases.
Charles' Law	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$	At constant P , as volume increases, temperature increases.
Gay-Lussac's Law	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$	At constant V , as pressure increases, temperature increases.
Combined Law	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$	Obtained by combining Boyle's Law, Charles' Law and Gay-Lussac's Law.
Ideal Gas Law	$PV = nRT$	
V = volume in dm^3 T = temperature in K		P = pressure in kPa n = number of moles R = ideal gas constant

Mixed Gas Law Calculation Work For Problems

Charles Henrickson



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