If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56 °C, what is the volume of the container that the gas is in?

$$V = \frac{nRT}{P} = \frac{7.7 \times 329 \times .0821}{.090tm} = 2310.9L$$

 If I have 17 moles of gas at a temperature of 67 °C, and a volume of 88.89 liters, what is the pressure of the gas?

 If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have

$$N = \frac{RT}{PV}$$
 $\frac{.0821 \times 300}{.5a4m \times 25L} = 1.97mol$

If I have 21 moles of gas held at a pressure of 78 atm and a temperature of 900 K, what is the volume of the gas?

10) If I have 1.9 moles of gas held at a pressure of 5 atm and in a container with a volume of 50 liters, what is the temperature of the gas?

11) If I have 2.4 moles of gas held at a temperature of 97.0°C and in a container with a volume of 45 liters, what is the pressure of the gas?370

Mixed Gas Laws Answers

Christopher R. Hren, John T. Moore, Peter J. Mikulecky

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