

## Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)



- a. If 15.0g of  $\text{N}_2\text{O}_4$  was produced, how many moles of  $\text{O}_2$  were required?

$$\frac{15.0\text{g N}_2\text{O}_4}{92.0\text{g N}_2\text{O}_4} \times \frac{1 \text{ mol N}_2\text{O}_4}{1 \text{ mol N}_2\text{O}_4} \times \frac{2 \text{ mol O}_2}{1 \text{ mol N}_2\text{O}_4} = 0.326 \text{ mol O}_2$$

- b. If  $4.0 \times 10^{-3}$  moles of oxygen reacted, how many grams of  $\text{N}_2$  were needed?

$$\frac{4.0 \times 10^{-3} \text{ mol O}_2}{2 \text{ mol O}_2} \times \frac{1 \text{ mol N}_2}{1 \text{ mol O}_2} \times \frac{28.0\text{g N}_2}{1 \text{ mol N}_2} = 5.6 \times 10^{-2} \text{ g N}_2$$



- a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

$$\frac{2.04 \text{ mol K}_3\text{PO}_4}{1 \text{ mol K}_3\text{PO}_4} \times \frac{3 \text{ mol KNO}_3}{1 \text{ mol K}_3\text{PO}_4} \times \frac{101.1 \text{ g KNO}_3}{1 \text{ mol KNO}_3} = 619\text{g KNO}_3$$

- b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

$$\frac{5.80\text{g AlPO}_4}{122.0\text{g AlPO}_4} \times \frac{1 \text{ mol AlPO}_4}{1 \text{ mol AlPO}_4} \times \frac{1 \text{ mol Al}(\text{NO}_3)_3}{1 \text{ mol AlPO}_4} = 4.75 \times 10^{-2} \text{ mol Al}(\text{NO}_3)_3$$



- a. If you have 5.50 mol of  $\text{CaC}_2$ , how much  $\text{C}_2\text{H}_2$  do you get?

$$\frac{5.50 \text{ mol CaC}_2}{1 \text{ mol CaC}_2} \times \frac{1 \text{ mol C}_2\text{H}_2}{1 \text{ mol CaC}_2} \times \frac{26.0\text{g C}_2\text{H}_2}{1 \text{ mol C}_2\text{H}_2} = 143\text{g C}_2\text{H}_2$$

- b. How many moles of water are needed when 65.0g of  $\text{CaC}_2$  have reacted?

$$\frac{65.0\text{g CaC}_2}{64.1\text{g CaC}_2} \times \frac{1 \text{ mol CaC}_2}{1 \text{ mol CaC}_2} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol CaC}_2} = 2.03 \text{ mol H}_2\text{O}$$

4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose ( $\text{C}_6\text{H}_{12}\text{O}_6$ ). Write and balance the chemical equation. How many moles of  $\text{CO}_2$  are required to make 120.0g of glucose?



$$\frac{120.0\text{g C}_6\text{H}_{12}\text{O}_6}{180.0 \text{ g C}_6\text{H}_{12}\text{O}_6} \times \frac{1 \text{ mol C}_6\text{H}_{12}\text{O}_6}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} \times \frac{6 \text{ mol CO}_2}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} = 4.000 \text{ mol CO}_2$$

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**Diana Harris**



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