

Unit 7: Balancing Chemical Reactions- Worksheet 2

Balance the following equations by inserting the proper coefficients.

- $\underline{2} \text{ SO}_2 + \underline{1} \text{ O}_2 \rightarrow \underline{2} \text{ SO}_3$ **Synthesis**
 - $\underline{2} \text{ CH}_4 + \underline{3} \text{ O}_2 \rightarrow \underline{2} \text{ CO} + \underline{4} \text{ H}_2\text{O}$ **Single R**
 - $\underline{2} \text{ P} + \underline{3} \text{ Cl}_2 \rightarrow \underline{2} \text{ PCl}_3$ **Synthesis**
 - $\underline{2} \text{ CO} + \underline{1} \text{ O}_2 \rightarrow \underline{2} \text{ CO}_2$ **Synthesis**
 - $\underline{2} \text{ CH}_4 + \underline{1} \text{ O}_2 \rightarrow \underline{2} \text{ CH}_3\text{OH}$ **Synthesis**
 - $\underline{2} \text{ Li} + \underline{1} \text{ Br}_2 \rightarrow \underline{2} \text{ LiBr}$ **Synthesis**
 - $\underline{2} \text{ Al}_2\text{O}_3 \rightarrow \underline{4} \text{ Al} + \underline{3} \text{ O}_2$ **Decomp**
 - $\underline{2} \text{ Na} + \underline{2} \text{ H}_2\text{O} \rightarrow \underline{2} \text{ NaOH} + \underline{1} \text{ H}_2$ **Single R**
 - $\underline{6} \text{ CO}_2 + \underline{6} \text{ H}_2\text{O} \rightarrow \underline{1} \text{ C}_6\text{H}_{12}\text{O}_6 + \underline{6} \text{ O}_2$ **Rev. Comb**
 - $\underline{1} \text{ H}_2\text{SO}_4 + \underline{2} \text{ NaCl} \rightarrow \underline{2} \text{ HCl} + \underline{1} \text{ Na}_2\text{SO}_4$ **Double R**
- Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:
- $\underline{3} \text{ H}_2 + \underline{1} \text{ SO}_2 \rightarrow \underline{1} \text{ H}_2\text{S} + \underline{2} \text{ H}_2\text{O}$ **Double R**
 - $\underline{2} \text{ CaCO}_3 + \underline{2} \text{ SO}_2 + \underline{1} \text{ O}_2 \rightarrow \underline{2} \text{ CaSO}_4 + \underline{2} \text{ CO}_2$ **No**
 - $\underline{2} \text{ AgNO}_3 + \underline{1} \text{ CaCl}_2 \rightarrow \underline{2} \text{ AgCl} + \underline{1} \text{ Ca(NO}_3)_2$ **Double R**
 - $\underline{2} \text{ HCl} + \underline{1} \text{ Ba(OH)}_2 \rightarrow \underline{1} \text{ BaCl}_2 + \underline{2} \text{ H}_2\text{O}$ **Double R**
 - $\underline{1} \text{ H}_3\text{PO}_4 + \underline{3} \text{ NaOH} \rightarrow \underline{1} \text{ Na}_3\text{PO}_4 + \underline{3} \text{ H}_2\text{O}$ **Double R**
 - $\underline{1} \text{ Pb(NO}_3)_2 + \underline{2} \text{ KI} \rightarrow \underline{1} \text{ PbI}_2 + \underline{2} \text{ KNO}_3$ **Double R**
 - $\underline{3} \text{ CuO} + \underline{2} \text{ NH}_3 \rightarrow \underline{1} \text{ N}_2 + \underline{3} \text{ Cu} + \underline{3} \text{ H}_2\text{O}$ **Single R**
 - $\underline{1} \text{ C}_2\text{H}_5\text{OH} + \underline{3} \text{ O}_2 \rightarrow \underline{2} \text{ CO}_2 + \underline{3} \text{ H}_2\text{O}$ **Com**
 - $\underline{2} \text{ C}_2\text{H}_6 + \underline{3} \text{ O}_2 \rightarrow \underline{2} \text{ CH}_3\text{COOH} + \underline{2} \text{ H}_2\text{O}$ **No**
 - $\underline{3} \text{ NO}_2 + \underline{1} \text{ H}_2\text{O} \rightarrow \underline{2} \text{ HNO}_3 + \underline{1} \text{ NO}$ **No**

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**Thirumalaisamy P.
Velavancorresponding**



Modeling Chemistry Ws Answers Balancing Chemical Reaction:

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