CHAPTER 14 REVIEW

Acids and Bases

SECTION 1

SHORT ANSWER Answer the following questions in the space provided.

1. Name the following compounds as acids:

sulfuric acid a. H₂SO₄

sulfurous acid b. H₂SO₃

hydrosulfuric acid e. H₂S

perchloric acid d. HClO₄

hydrocyanic acid e. hydrogen cyanide

- 2. H₂S Which (if any) of the acids mentioned in item 1 are binary acids?
- Write formulas for the following acids:

HNO₂ a. nitrous acid

HBr b, hydrobromic acid

H₃PO₄ c. phosphoric acid

CH₃COOH d. acetic acid

HCIO e. hypochlorous acid

Calcium selenate has the formula CaSeO₄.

H₂SeO₄ a. What is the formula for selenic acid?

H₂SeO₃ b. What is the formula for selenous acid?

Use an activity series to identify two metals that will not generate hydrogen gas when treated with an acid.

Choose from Cu, Ag, Au, Pt, Pd, or Hg.

6. Write balanced chemical equations for the following reactions of acids and bases:

a. aluminum metal with dilute nitric acid

 $2AI(s) + 6HNO_3(aq) \rightarrow 2AI(NO_3)_3(aq) + 3H_2(g)$

b. calcium hydroxide solution with acetic acid

 $Ca(OH)_2(aq) + 2CH_3COOH(aq) \rightarrow Ca(CH_3COO)_2(aq) + 2H_2O(I)$

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