

CLASS SET

Mixed Stoichiometry Problems

1. How many moles of H_2 would be required to completely react with O_2 to produce 5 moles of water? **5 mol H_2**
2. $H_2SO_4 + NaOH \rightarrow Na_2SO_4 + H_2O$
 - a. Balance this equation
 - b. What mass of H_2SO_4 would be required to react with 0.75 mol of NaOH?
37g
3. What mass of NO_2 is formed when NO reacts with 384 g of O_2 ? **1104g**
4. If 30 g of H_2SO_4 reacts with aluminum hydroxide in a double replacement reaction, what mass of water is produced? **11.0g**
5. If 10 L of CH_4 combusts with O_2 , what volume of water will be produced? **20 L H_2O**
6. Tin(II) fluoride, SnF_2 , is used in some toothpastes. It is made by the reaction of tin with hydrogen fluoride according to the following equation.
 $Sn + 2HF \rightarrow SnF_2 + H_2$
How many grams of SnF_2 are produced from the reaction of 30 g of HF with Sn? **118 g**
7. In a spacecraft, the carbon dioxide exhaled by astronauts can be removed by its reaction with lithium hydroxide, $LiOH$, according to the following chemical equation.
 $CO_2(g) + 2LiOH(s) \rightarrow Li_2CO_3(s) + H_2O(l)$
If 448 L of CO_2 is exhaled, the average amount exhaled by a person each day, how much (in grams) $Li_2CO_3(s)$ is produced? **1480g Li_2CO_3**
8. When 9.8g of aluminum oxide decomposes, how many grams of aluminum metal are produced? **5.2 g Al**
9. How many grams of iodine are produced when 0.72 mol of fluorine react with potassium iodide?
10. How many grams of sodium are required to react with water to produce 5.0g of sodium hydroxide? ($Na + H_2O \rightarrow NaOH + H_2$)
11. How many grams of H_2 are produced in the reaction in #10?
12. What volume would be occupied by the resulting hydrogen at STP?
13. 20.0g of K react with H_2O to produce KOH and H_2 gas. How many liters would the hydrogen gas occupy at STP?
14. If 30.2g of Al react with HCl to produce $AlCl_3$ and H_2 gas, how many liters of H_2 are produced at STP?
15. How many liters of HCl are produced by the reaction of 5.7L of H_2 with Cl_2 ?

Mixed Stoichiometry Word Problems And Answers

TD Snyder



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