

Formula Stoichiometry Practice. *This sheet should be printed, completed by hand, and scanned to produce a single pdf file with filename "lastname_FormulaPract" for upload to the dropbox.*

Name: _____

1. Write down the indicated mole ratios for potassium chlorate, KClO_3 . The numerator is listed first in each case.

moles K to moles O

$$\frac{1 \text{ mol K}}{3 \text{ mol O}}$$

moles K to moles Cl

$$\frac{1 \text{ mol K}}{1 \text{ mol Cl}}$$

moles K^+ to moles ClO_3^-

$$\frac{1 \text{ mol K}}{1 \text{ mol ClO}_3^-}$$

moles Cl to O

$$\frac{1 \text{ mol Cl}}{3 \text{ mol O}}$$

moles Cl to moles K

$$\frac{1 \text{ mol Cl}}{1 \text{ mol K}}$$

moles O to moles K

$$\frac{3 \text{ mol O}}{1 \text{ mol K}}$$

moles O to moles Cl

$$\frac{3 \text{ mol O}}{1 \text{ mol Cl}}$$

moles ClO_3^- to moles K^+

$$\frac{1 \text{ mol ClO}_3^-}{1 \text{ mol K}^+}$$

2. Calculate molar masses for the two compounds below making sure to show your work.

potassium chlorate, KClO_3

$$\begin{array}{r} \text{K} = 39.098 \\ \text{Cl} = 35.45 \\ + \text{O} = 15.999 \times 3 = 47.997 \\ \hline 122.545 \text{ g/mol} \end{array}$$

aluminum sulfate, $\text{Al}_2(\text{SO}_4)_3$

$$\begin{array}{r} \text{Al} = 26.982 \times 2 = 53.964 \\ \text{SO}_4 = [(15.999 \times 4) + 32.06] \times 3 = 288.168 \\ \hline 342.132 \text{ g/mol} \end{array}$$

Mixed Stoich Practice

AW Chickering



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