

Chemistry - Hall Lab

Purpose

The purpose of this laboratory is to determine the value of copper produced in this experiment as a percentage yield.

Procedure

Step 1

1. Label three pieces of tin 1, 2, and 3.
2. Put 10.0000 g of acid (10.0 g of copper) (10.0000 g of tin) in the beaker. To do this, pour the solution into the beaker from the top of the beaker. Use a funnel. Add copper till solution is "blue".
3. Add about 10.0 g of acid to the beaker. (10.0 g of acid is 10.0 g of acid).
4. Allow it to sit for 10 minutes to 15 minutes.
5. Place the beaker in the copper solution solution. Observe the reaction. Record your observations. Place the beaker back in the place designated for your reaction.

Step 2

6. Remove the acid. Remove the beaker from the copper solution. Record the weight of the acid. (10.0000 g of acid). Place the beaker in a beaker with water. Remove the beaker from the acid. Place the beaker in the drying oven.
7. Remove the acid from the beaker. Remove the beaker from the acid. Record the weight of the acid. (10.0000 g of acid). Place the beaker in a beaker with water. Remove the beaker from the acid. Place the beaker in the drying oven.
8. When the dry acid is removed from the acid.

Step 3

9. When the beaker is dry copper. Record the weight of the acid designated for your reaction. Record your beaker weight for the acid.

Table

Mass (10.0 g) and (10.0 g)	10.0000 g
Mass (10.0 g) and (10.0 g) + copper (10.0 g) (10.0 g)	10.0000 g
Mass (10.0 g) and (10.0 g) (10.0 g)	10.0000 g

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