#### Chemistry Unit 7 Review

#### 701 - Inquiry Lab

- Define...
  - a. Direct relationship: Two variables increase/decrease together
  - b. Indirect relationship: One variable increases, the other decreases
- 2. Identify whether the listed variables have a direct or indirect relationship to each other.
  - a. Moles and volume: Direct

- c. Pressure and volume: Indirect
- Temperature and pressure; Direct
- d. Moles and pressure: Direct
- 3. Using the Ideal Gas Law equation, list 3 ways you could increase the volume of a balloon.

Increase the moles, increase the temperature, decrease the pressure (moreso, the pressure drops as a result of greater volume)

4. Using the Ideal Gas Law equation, list 3 ways you could decrease the pressure within a syringe.

Decrease the moles, decrease the temperature, increase the volume

5. If the number of moles in a container is doubled, by how much should the pressure increase (assuming everything else is held constant)?

The pressure should also double

#### 702 - Gas Laws Equation

1. What equation should you use when you are given just one set of variables?

PV = nRT

 You have a 4.5 L canister of O<sub>2</sub> that you store in a room that is 20 °C. The pressure of this tank is 13000 kPa. How many moles of O<sub>2</sub> are in this canister?

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**Daniel F McAuley** 

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