

Chemistry - Hall Lab

Purpose

The purpose of this laboratory is to determine the value of copper produced in this experiment as a percentage yield.

Procedure

Step 1

1. Label three pieces of 100 mL beakers.
2. Put 10.0000 g of solid 100% Zn in beaker 1, 10.0000 g of aluminum in beaker 2, and 10.0000 g of copper in beaker 3. Label each beaker up to the other two beakers. Label each beaker with the amount of Zn, Al, and Cu.
3. Add 100 mL of 1.0 M HCl to each beaker. Stir with a glass rod.
4. Allow 10 min for the reaction to take place.
5. Place the beakers in the fume hood. Observe the reactions. Record your observations. Place the beakers in the fume hood designated for your reaction.

Step 2

6. Remove the beakers. Remove the beakers from the fume hood. Record the mass of the beakers. Place the beakers in a fume hood designated for your reaction. Record the appearance of the beakers. Place the beakers in the fume hood.
7. Remove the beakers from the fume hood. Remove the beakers from the fume hood. Record the mass of the beakers. Place the beakers in a fume hood designated for your reaction. Record the appearance of the beakers. Place the beakers in the fume hood.
8. Remove the beakers from the fume hood. Remove the beakers from the fume hood. Record the mass of the beakers. Place the beakers in a fume hood designated for your reaction. Record the appearance of the beakers. Place the beakers in the fume hood.

Step 3

9. Remove the beakers from the fume hood. Remove the beakers from the fume hood. Record the mass of the beakers. Place the beakers in a fume hood designated for your reaction. Record the appearance of the beakers. Place the beakers in the fume hood.

Results

Mass of Zn and Al	10.0000 g
Mass of Zn and Al + copper (10.0000 g)	10.0000 g
Mass of Zn and Al + copper (10.0000 g)	10.0000 g

Modeling Chemistry Nail Lab

J Dewey



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