

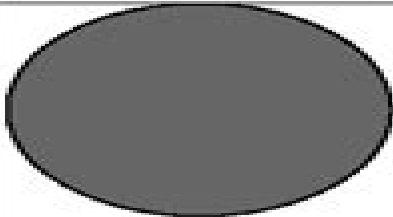
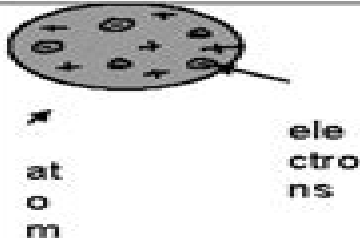

# Atomic Theory Worksheet

*Objective: You will create an abbreviated atomic theory timeline to show how the scientists involved in the atomic theory worked together to construct the modern model of the atom.*

## Part 1 – Atomic Theory Timeline

Choose 3 scientists from the options below and complete the table.

Options: John Dalton, Niels Bohr, JJ Thomson, Erwin Schrödinger, Ernest Rutherford

Content	Scientist 1	Scientist 2	Scientist 3
Timeframe [in order]	1803	1897	1911
Scientist Name	John Dalton	J.J. Thompson	Ernest Rutherford
What did he contribute to the atomic theory?	He contributed that atoms can not be made or destroyed, a chemical reaction only can happen when atoms are rearranged, as well all atoms of a specific element are the same in mass and properties.	He contributed that atoms have charged particles that are called electrons, he also thought that matter is electronically neutral, he also thought that electrons are dispersed evenly through an atom.	He contributed that an atom has a very small positively charged center called the nucleus and it contains most of the atoms mass. He also added that atoms have a cloud of electrons that circle around the nucleus.
Draw a sketch of their model [by hand or using the drawing tool] and label all parts.		 atom electrons	 electrons nucleus

## Part 2 – Reflection Questions

1. Describe two details of the modern atomic model.

Atoms have a nucleus of protons and neutrons and a cloud around the nucleus that contains the electrons.

2. Explain how the atomic model has changed over time.

The Rutherford model got replaced by the Bohr model because scientists found out that electrons have a specific orbit around the nucleus.

# Modern Atomic Theory Concept Review Answers

**M Mark**



## **Modern Atomic Theory Concept Review Answers:**

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**Foundations of College Chemistry** Morris Hein, Susan Arena, Cary Willard, 2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students. It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence. The central focus is to make chemistry interesting and understandable and teach students the problem-solving skills they will need. This International Adaptation offers new and updated content with improved presentation of all course material. It builds on the strengths of previous editions including clear explanations and step-by-step problem solving. The material emphasizes real-world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry. There is new and expanded coverage of polarizing power and polarizability, Fajans' rules, collision number and mean free path, abnormal molecular masses and van't Hoff factor and applications of radioactivity. **Study Guide to accompany Basic Concepts of Chemistry, 7th Edition** Leo J. Malone, 2003-02-20

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### **Table of Contents Modern Atomic Theory Concept Review Answers**

1. Understanding the eBook Modern Atomic Theory Concept Review Answers
  - The Rise of Digital Reading Modern Atomic Theory Concept Review Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Atomic Theory Concept Review Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Atomic Theory Concept Review Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Atomic Theory Concept Review Answers
  - Personalized Recommendations
  - Modern Atomic Theory Concept Review Answers User Reviews and Ratings
  - Modern Atomic Theory Concept Review Answers and Bestseller Lists

5. Accessing Modern Atomic Theory Concept Review Answers Free and Paid eBooks
  - Modern Atomic Theory Concept Review Answers Public Domain eBooks
  - Modern Atomic Theory Concept Review Answers eBook Subscription Services
  - Modern Atomic Theory Concept Review Answers Budget-Friendly Options
6. Navigating Modern Atomic Theory Concept Review Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Modern Atomic Theory Concept Review Answers Compatibility with Devices
  - Modern Atomic Theory Concept Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modern Atomic Theory Concept Review Answers
  - Highlighting and Note-Taking Modern Atomic Theory Concept Review Answers
  - Interactive Elements Modern Atomic Theory Concept Review Answers
8. Staying Engaged with Modern Atomic Theory Concept Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Atomic Theory Concept Review Answers
9. Balancing eBooks and Physical Books Modern Atomic Theory Concept Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Atomic Theory Concept Review Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modern Atomic Theory Concept Review Answers
  - Setting Reading Goals Modern Atomic Theory Concept Review Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Atomic Theory Concept Review Answers
  - Fact-Checking eBook Content of Modern Atomic Theory Concept Review Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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