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 $\text{NaCl} + \text{F}_2 \rightarrow \text{NaF} + \text{Cl}_2$
 $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
 $\text{Pb} + \text{OH}^- \rightarrow \text{PbCl}_2 + \text{H}_2\text{O}$
 $2\text{Al} + 3\text{K}_2\text{SO}_4 \rightarrow 6\text{KBr} + 2\text{Al}_2\text{O}_3$
 $4\text{CH}_4 + 12\text{O}_2 \rightarrow 4\text{CO}_2 + 8\text{H}_2\text{O}$
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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
 $2\text{C}_6\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr}_3 + \text{NaBr} \rightarrow \text{NaNO}_3 + \text{Ga}_2\text{O}_3$
 $4\text{NaNO}_3 \rightarrow 2\text{N}_2 + 3\text{O}_2 + 2\text{Na}_2\text{O}$
 $3\text{NF}_3 + 5\text{H}_2\text{O} \rightarrow 3\text{HNO}_2 + 9\text{HF}$
 $2\text{N}_2\text{O}_3 \rightarrow 4\text{NO}_2 + \text{O}_2$ using the equation from problem 2 above answer the following

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 $2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$
 $3\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 \rightarrow 2\text{Ga}_2\text{O}_3 + 3\text{SO}_2 + 6\text{NaBr}$
 $3\text{NaNO}_3 \rightarrow 2\text{N}_2 + 3\text{O}_2 + 2\text{Na}_2\text{O}$
 $3\text{NF}_3 + 5\text{H}_2\text{O} \rightarrow 3\text{HNO}_2 + 9\text{HF}$
 $2\text{N}_2\text{O}_3 \rightarrow 4\text{NO}_2 + \text{O}_2$ using the

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$\text{ZnCl}_2 + \text{Cu} + \text{Zn} + \text{CuCl}_2 \rightarrow 2 \text{ZnCl}_2 + 2 \text{Cu}$ how many moles of $\text{Ce} + \text{ZnCl}_2 + \text{ZnCl}_2$ will be produced from 23.0 g of $\text{Ce} + \text{Zn}$

assuming $\text{Ce} + \text{CuCl}_2 + \text{CuCl}_2$ is available in excess

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web jul 22 2013 balance the following equations
 $\text{Br}_2 + \text{H}_2\text{O} \rightarrow \text{HBr} + \text{HCO}_3^-$
 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + \text{NaBr}$
 $\text{SNO}_2 + \text{NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3$ using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

$\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2 \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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web the stoichiometry of a balanced chemical equation identifies the maximum amount of product that can be obtained the

stoichiometry of a reaction describes the relative amounts of reactants and products in a balanced chemical equation

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equations
 $\text{N}_2 + \text{F}_2 + \text{NF}_3 + \text{C}_6\text{H}_{10} + \text{O}_2 + \text{CO} + \text{H}_2\text{O} + \text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 + \text{GaBr} + \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + \text{NaBr} + \text{SNO} + \text{NF}_3 + \text{SnF}_2$

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web may 28 2020 write the balanced equation for the reaction and calculate how much SiO_2 is required to produce 3.00 kg of SiC
 $\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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equations
 $\text{N}_2 + \text{F}_2 + \text{NF}_3 + \text{C}_6\text{H}_{10} + \text{O}_2 + \text{CO} + \text{H}_2\text{O} + \text{HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{GaBr} + \text{Na}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + \text{NaBr} + \text{SNO} + \text{NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3 + \text{CO}_2 + \text{NaBr}$ solve the following stoichiometry grams grams problems 6

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answers

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web gabrielle m 9 years ago i m working on chemical reactions double and single replacement on flvs now my first question for this video is how do you have 4 aluminum atoms when it says 2Al subscript 2 do you just add the 2 and the subscript 2 same thing goes with the o3 i am so lost

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web balancing equations and simple stoichiometry answers balance the following equations $1\text{ N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$ $2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$ $3\text{HBr} + 1\text{KHCO}_3 \rightarrow 1\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2$ $4\text{Ga} + 3\text{Na}_2\text{SO}_3 \rightarrow 2\text{Ga}_2\text{SO}_3 + 3\text{Na}_2\text{S}$ $5\text{Sn} + 2\text{NF}_3 \rightarrow 3\text{SnF}_2 + 1\text{N}_2\text{O}_3$

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web google classroom balance the following chemical equation $\text{Mg} + \text{OH}^- \rightarrow \text{Mg}(\text{OH})_2$ $\text{HCl} + \text{MgCl}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}$ note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem $7.4 \times 10^5 \theta \pi 8.5 \times 10^9$

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web a common type of stoichiometric relationship is the mole ratio which relates the amounts in moles of any two substances in a chemical reaction we can write a mole ratio for a pair of substances by looking at the coefficients in front of each species in

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