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# Jp University Chapra Program Physics 2013

**Camilla Rothe**



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web what is a titration a titration is an analytical procedure used to determine the accurate concentration of a sample by reacting it with a standard solution one type of titration uses a neutralization reaction in which an acid and a base react to produce a salt and water

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web donor whereas a base generically represented by b is a proton acceptor when acids and bases combine they swap a proton ha b a hb strong acids e g sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web jul 19 2023 at the beginning of the titration the solution in the erlenmeyer flask is acidic as the base is added it completely reacts with the acid and the solution in the erlenmeyer flask continues to be acidic but at the equivalence point acid base titration chemistry 1210 lab report containing an - Oct 07 2023

web answer based on the data the strong acid strong base combination of hydrochloric acid and sodium hydroxide had the neutral ph at the equivalence point being 7 this result is expected because a combination of strong acids and strong bases should result in a neutral at the equivalence point

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web 1 approximately 3 g of hydrated oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  was weighted to the nearest 0 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of  $\text{HCl}$  because this is a strong acid the ionization is complete and the hydronium ion molarity is 0 100 m the ph of the solution is then  $\text{pH} = \log 0 100 = 1 000$   $\text{pH} = \log 0 100 = 1 000$  b titrant volume 12 50 ml

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web one type of titration uses a neutralization reaction in which an acid and a base react to produce a salt and water in equation 1 the acid is  $\text{HCl}$  hydrochloric acid and the base is  $\text{NaOH}$  sodium hydroxide when the acid and base react they form  $\text{NaCl}$  sodium chloride which is also known as table salt

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