

Question 2

2(a)(i) Comparison: X is more easily ionized than Mg.

Explanation: X is lower in group II than Mg. The outer electrons of X are in a shell further away from the nucleus than those of Mg. The outer electrons of X are therefore less strongly held by the nucleus and more easily lost than the outer electrons of Mg.

2(a)(ii) Comparison: Cl is a stronger oxidizing agent than Y.

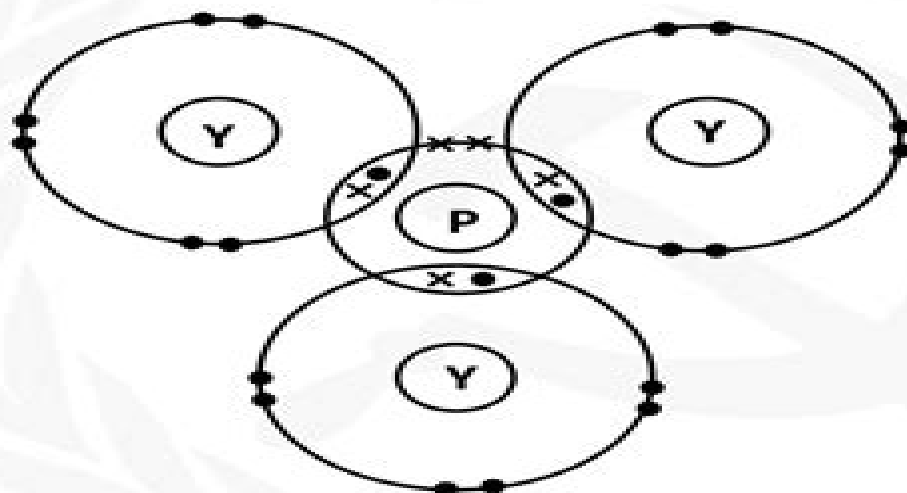
Explanation: Oxidizing power of elements decreases down group VII. Cl is higher in group VII than Y. The nucleus of Cl is therefore less shielded than that of Y, making external electrons more strongly attracted to the nucleus of Cl than of Y.

2(b)(i) Both elements are solid at R.T.P.

2(b)(ii) At R.T.P. X conducts electricity but Y does not.

2(c)(i) Covalent bonding occurs between phosphorous and element Y.

200000



Key

● = Y electron

 $\kappa = \rho \text{ electron}$

Y = nucleus and inner electrons of Y.

(P) = nucleus and inner electrons of P

2(c)(iii) Chemical formula: Py_n

Two properties of compound:

1. It is a solid with a low melting point.
2. It is soluble in non-polar (organic) solvents.

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