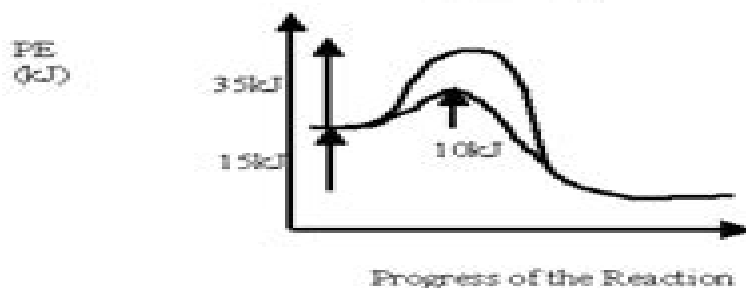


1. Which of the following units could be used to express the reaction rate?
- A. mL/s
B. mL/g
C. g/mL
D. mL/mol
2. Consider the reaction: $\text{Zn}_{(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{ZnCl}_{2(aq)} + \text{H}_{2(g)}$
The rate of production of ZnCl_2 can be increased by
- A. decreasing the $[\text{HCl}]$.
B. increasing the temperature.
C. increasing the volume of H_2 .
D. decreasing the surface area of Zn.
3. The statement, the minimum energy needed for a successful collision, defines
- A. enthalpy.
B. activation energy.
C. the ΔH of the reaction.
D. the activated complex.
4. As an activated complex changes to products,
- A. potential energy changes to kinetic energy.
B. kinetic energy changes to potential energy.
C. kinetic energy changes to activation energy.
D. potential energy changes to activation energy.
5. Which of the following is most likely to have the greatest rate at room temperature.
- A. $2\text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2\text{H}_2\text{O}_{(l)}$
B. $2\text{Ag}_{(s)} + \text{CrO}_4^{2-}{}_{(aq)} \rightarrow \text{Ag}_2\text{CrO}_{4(s)}$
C. $\text{Pb}_{(l)} + 2\text{HCl}_{(aq)} \rightarrow \text{PbCl}_{2(aq)} + \text{H}_{2(g)}$
D. $\text{CH}_{4(g)} + 2\text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$
6. Consider the following PE diagram for an uncatalyzed and catalyzed reaction



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L Towner



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