

Energie basse : je vais
m'effondrer
(me cacher, perdre espoir)

COMMENT JE ME SENS

Emotions débordantes : je vais
exploser
(casser, taper, crier)

Zone bleue

Comment je me sens

Triste



Seul



Fatigué



Malade



Ce que je peux faire

**Me reposer,
demander du
soutien**

Je dors, je fais la sieste.

Je m'étire.

Je parle à
quelqu'un.

Je peux aussi

Zone verte

Comment je me sens

Content



Confiant



Calm



Concentré



Ce que je peux faire

**Continuer mes
projets, apprendre,
apporter de l'aide**

Je fais une activité créative.

Je finis un travail

Je m'entraîne.

J'écoute un ami qui ne
va pas bien.

Je peux aussi

Zone orange

Comment je me sens

Indécis



Contrarié



Impuissant



Soucieux



Ce que je peux faire

**Ralentir, me
recentrer, me
faire du bien**

Je masse mes tempes,
mes yeux, mes pieds.

Je vais marcher.

Je prends l'air.

Je peux aussi

Zone rouge

Comment je me sens

Frustré



Paniqué



Furieux



Submergé



Ce que je peux faire

**Faire une pause,
respirer, me faire
aider**

J'expire lentement par la
bouche.

Je malaxe une balle.

Je bois de l'eau fraîche.

Je peux aussi

Je Me Sens En Deacuteecalage Et Je Reacuteeussis Ma Vie

Tobias Bleicker



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baris kalimat yang disebut dengan gatra ada pula aturan bait satu dengan bait lainnya yang harus saling memiliki keterkaitan suku kata yang disebut guru lagu

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web 1 define each remember oil rig oxidation is loss of e reduction is gain of e a oxidation b reduction c oxidizing agent d reducing agent 2 write half reactions for each of the following atoms or ions label each as oxidation or reduction use the activity series chart a al b ba2 c br 2 d ca e ga3 f h 2 g h

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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web you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 identify the oxidation state of each element in the following a na 2 so 4 b mno 4 c kclo 3 d ch 3 oh 2 identify the oxidation state of nitrogen in the following a no 3 b no 2 c no 2

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web an oxidation reduction redox reaction is one in which one or more atoms change oxidation numbers oxidation occurs when an atom s oxidation state becomes more positive indicating that electrons have been lost reduction occurs when an atom s oxidation state becomes less positive indicating that electrons have been gained

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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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web what is an oxidation reduction or redox reaction this is a reaction in which electrons are transferred between reactants what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is a helpful way to remember this

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23 fe²⁺ co co₂ fe

oxidation co co₂ 2e reduction fe²⁺ 2e fe 24 3 ag ni₃ 3 ag oxidation reduction 25 cu₂ pb pb₂ cu oxidation reduction 26 o 2

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a $i_2 + o_5 \rightarrow 5co + i_2 + o_5$

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also $2sr + o_2 \rightarrow 2sro$ sr^0 to sr^2 oxidized reducing agent o^0 to o^2 reduced ox ag $2li + s \rightarrow li_2s$ li^0 to li^1 oxidized red ag s^0 to s^2 reduced ox ag $3 2cs + br_2$