

06. The unit of K_c for reaction $N_2 + O_2 \rightleftharpoons 2NO$

- (a) mol dm^{-3} (b) $\text{mol}^{-2} \text{dm}^6$
(c) $\text{mol}^{-1} \text{dm}^3$ (d) ☒ no unit

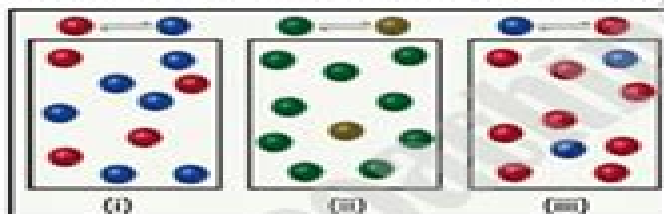
07. The system is stable in equilibrium when:

- (a) ☒ $Q_c = K_c$ (b) $Q_c > K_c$
(c) $Q_c < K_c$ (d) None of these

08. Q_c can be defined as:

- (a) ratio of product and reactants
(b) ☒ ratio of molar concentration of product and reactant at specific time
(c) ratio of molar concentration of product and molar volume of reactant
(d) ratio of molar concentration of product and reactant raised to the power of coefficient

09. Which of the following represent backward reaction?



- (a) (i) and (ii) (b) (ii) and (iii)
(c) (ii) only (d) ☒ (iii) only

10. The value of K_c increases when:

- (a) [Product] is less (b) ☒ [Product] is more
(c) [Reactant] is more (d) [Reactant = product]

11. Which of these are reversible reactions?

- (a) Condensation (b) Evaporation
(c) ☒ Condensation & Evaporation (d) None of them

Multiple Choice Questions In Chemical Equilibrium

Bloomfield, Catherine E. MacGowan



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