

Chemistry Unit 6 Worksheet 1: Intermolecular Forces

1. Predict the molecular shape of each of the following:

- a. H_2S _____ b. CCl_4 _____
 c. SO_2 _____ d. PCl_5 _____

2. List all types of IMFs that would occur in each of the following (you should have a good enough understanding of electronegativities to answer all of these, except maybe SO_2 , without look up the electronegativity numbers).

- a. CH_3CF_3 _____
 b. CCl_4 _____
 c. SO_2 _____
 d. BrF _____
 e. $(\text{CH}_3)_3\text{N}$ _____
 f. PCl_5 _____

3. H_2S , O_2 and CH_3OH all have comparable molecular masses. List the dominant type of IMF for the pure substances, then rank the strength of each compound based on IMFs within the samples.

4. Circle all of the species below that can form a hydrogen bond in its pure form. Explain why the other species couldn't hydrogen bond.



5. Rank the following compounds from weakest intermolecular forces to strongest. Justify your answers.



6. Rank the following from weakest intermolecular forces to strongest. Justify your answers.



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Barbara Ucko



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