63. The metal that cannot be obtained by electrolysis of an aqueous solution of its salts is :
(1) Cu (2) Cr (3) Ag (4) Ca

63. (4)
Factual

64. Resistance of 0.2 M solution of an electrolyte is 50 Ω. The specific conductance of the solution is 1.4 S m^{-1}. The resistance of 0.5 M solution of the same electrolyte is 280 Ω. The molar conductivity of 0.5 M solution of the electrolyte in S m^2 mol^{-1} is :
(1) 5×10^3 (2) 5×10^2 (3) 5×10^{-4} (4) 5×10^{-3}

64. (3)
$$\Lambda_m = \frac{k(s/m)}{1000 \times Molarity(mol/L)} = \frac{1}{4 \times 1000 \times 0.5}$$
$$= 5 \times 10^{-4} \, Sm^2 mol^{-1}$$

65. CsCl crystallises in body centred cubic lattice. If "a" is its edge length then which of the following expressions is correct ?
(1) $r_{Cs^+} + r_{Cl^-} = \frac{\sqrt{3}}{2}a$ (2) $r_{Cs^+} + r_{Cl^-} = \sqrt{3}a$ (3) $r_{Cs^+} + r_{Cl^-} = 3a$ (4) $r_{Cs^+} + r_{Cl^-} = \frac{3a}{2}$

65. (1)
Factual

66. Consider separate solutions of 0.500 M C$_2$H$_5$OH(aq), 0.100 M Mg$_3$(PO$_4$)$_2$(aq), 0.250 M KBr(aq) and 0.125 M Na$_3$PO$_4$(aq) at 25°C. Which statement is true about these solutions, assuming all salts to be strong electrolytes ?
(1) 0.125 M Na$_3$PO$_4$(aq) has the highest osmotic pressure.
(2) 0.500 M C$_2$H$_5$OH(aq) has the highest osmotic pressure.
(3) They all have the same osmotic pressure.
(4) 0.100 M Mg$_3$(PO$_4$)$_2$(aq) has the highest osmotic pressure.

66. (3)
All solutions have same osmotic pressure.

67. In which of the following reactions H$_2$O$_2$ acts as a reducing agent ?
(a) H$_2$O$_2$ + 2H$^+$ + 2e$^-$ → 2H$_2$O (b) H$_2$O$_2$ - 2e$^-$ → O$_2$ + 2H$^+$
(c) H$_2$O$_2$ + 2e$^-$ → 2OH$^-$ (d) H$_2$O$_2$ + 2OH$^-$ - 2e$^-$ → O$_2$ + 2H$_2$O
(1) (a), (c) (2) (b), (d) (3) (a), (b) (4) (c), (d)

67. (2)
H$_2$O$_2$ → O$_2$ + 2H$^+$ + 2e$^-$
H$_2$O$_2$ + 2OH$^-$ → O$_2$ + 2H$_2$O + 2e$^-$

68. In S$_N$2 reactions, the correct order of reactivity for the following compounds :
CH$_3$Cl, CH$_3$CH$_2$Cl, (CH$_3$)$_2$CHCl and (CH$_3$)$_3$CCl is :
(1) CH$_3$CH$_2$Cl > CH$_3$Cl > (CH$_3$)$_2$CHCl > (CH$_3$)$_3$CCl
(2) (CH$_3$)$_2$CHCl > CH$_3$CH$_2$Cl > CH$_3$Cl > (CH$_3$)$_3$CCl
(3) CH$_3$Cl > (CH$_3$)$_2$CHCl > CH$_3$CH$_2$Cl > (CH$_3$)$_3$CCl
(4) CH$_3$Cl > CH$_3$CH$_2$Cl > (CH$_3$)$_2$CHCl > (CH$_3$)$_3$CCl

68. (4)
Order of reactivity of alkyl halide towards SN-2 reaction is CH$_3$X > 1° > 2° > 3°

69. The octahedral complex of a metal ion M^{3+} with four monodentate ligands L$_1$, L$_2$, L$_3$ and L$_4$ absorb wavelengths in the region of red, green, yellow and blue, respectively. The increasing order of ligand strength of the four ligands is :
(1) L$_3$ < L$_2$ < L$_4$ < L$_1$ (2) L$_1$ < L$_2$ < L$_4$ < L$_3$ (3) L$_4$ < L$_3$ < L$_2$ < L$_1$ (4) L$_1$ < L$_3$ < L$_2$ < L$_4$

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