

Thermochemistry

Multiple Choice Questions

- B** 1. When the three types of energy changes are arranged in order of decreasing molar heat values, what is the correct order?
A. chemical, phase, nuclear **C.** nuclear, phase, chemical
B. nuclear, chemical, phase **D.** phase, chemical, nuclear
- B** 2. The temperature remains constant as energy is added to a substance. How may the substance be changing?
A. from a gas to a solid **C.** from a liquid to a solid
B. from a liquid to a gas **D.** in the amount of kinetic energy
- A** 3. Which metal requires the most energy to raise 1.00 g of it by 1.00°C?
A. aluminum **B.** copper **C.** lead **D.** nickel
- B** 4. A substance has a molar heat of combustion of -850.4 kJ/mol. When 0.285 mol of the substance is burned in a calorimeter containing 8.60 kg of water, what is the increase of the water temperature?
A. 0.156°C **B.** 6.41°C **C.** 7.89°C **D.** 12.8°C
- B** 5. The addition of 9.54 kJ of heat is required to raise the temperature of 225.0 g of a liquid hydrocarbon from 20.5°C to 45.0°C. What is the heat capacity of this hydrocarbon?
A. 0.94 J/g·°C **B.** 1.73 J/g·°C **C.** 1.88 J/g·°C **D.** 9.42 J/g·°C
- C** 6. If the molar heat of fusion of sodium is 2.63 kJ/mol, how much energy is needed to melt 180.0 g of solid sodium at its melting point?
A. 2.63 kJ **B.** 2.93 kJ **C.** 20.6 kJ **D.** 473 kJ
- B** 7. Which process is exothermic?
A. the boiling of liquid nitrogen **C.** the sublimation of dry ice
B. the freezing of water **D.** the vaporization of water
- A** 8. If the heat of formation for the chemical reaction, $\text{S}_8(\text{s}) + 12\text{O}_2(\text{g}) \rightarrow 8\text{SO}_2(\text{g})$, is -395.7 kJ/mol of SO_2 , how can the equation be rewritten to include the heat term?
A. $\text{S}_8(\text{s}) + 12\text{O}_2(\text{g}) \rightarrow 8\text{SO}_2(\text{g}) + 3166 \text{ kJ}$ **C.** $\text{S}_8(\text{s}) + 12\text{O}_2(\text{g}) \rightarrow 8\text{SO}_2(\text{g}) + 395.7 \text{ kJ}$
B. $\text{S}_8(\text{s}) + 12\text{O}_2(\text{g}) \rightarrow 8\text{SO}_2(\text{g}) - 395.7 \text{ kJ}$ **D.** $\text{S}_8(\text{s}) + 12\text{O}_2(\text{g}) + 3166 \text{ kJ} \rightarrow 8\text{SO}_2(\text{g})$
- A** 9. Which statement correctly describes the energy changes that take place when a solid changes to a liquid at constant temperature?
A. The potential energy increases, and the kinetic energy remains constant.
B. The potential energy remains constant, and the kinetic energy increases.
C. The potential energy decreases, and the kinetic energy decreases.
D. The potential energy increases, and the kinetic energy decreases.
- B** 10. Why does warm water have a more rapid rate of evaporation than cold water?
A. It has a higher viscosity.
B. Its molecules have a higher average kinetic energy.
C. Its molecules have more attraction for one another.
D. More of its molecules have nearly the same kinetic energy.
- B** 11. If the heat of fusion of a substance is 20 kJ/mol, what heat is released when 1.0 mol of liquid at the melting point freezes?
A. 10 kJ **B.** 20 kJ **C.** 40 kJ **D.** 60 kJ
- C** 12. The amount of 2.00 mol of a solid ionic compound was dissolved in 225 mL of water in an insulated container. The initial temperature of the water was 20.0°C and the final temperature of the water was 47.8°C. What is ΔH_{soln} ?
A. -27.8 kJ/mol **B.** -26.2 kJ/mol **C.** -13.1 kJ/mol **D.** -6.26 kJ/mol
- D** 13. What does it mean if the ΔH value for a chemical reaction is positive?
A. kinetic energy is increasing in the system
B. potential energy is decreasing in the system
C. products have less potential energy than reactants
D. reactants have less potential energy than products.
- A** 14. Living plants produce glucose in the process of photosynthesis according to this equation:
 $6\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{l}) + \text{energy} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6(\text{s}) + 6\text{O}_2(\text{g})$
 Is this reaction endothermic or exothermic, and is the value of ΔH° positive or negative?
A. endothermic, positive **B.** endothermic, negative **C.** exothermic, positive **D.** exothermic, negative

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JS Bruner



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Table of Contents Multiple Choice Questions And Answers On Thermochemistry

1. Understanding the eBook Multiple Choice Questions And Answers On Thermochemistry
 - The Rise of Digital Reading Multiple Choice Questions And Answers On Thermochemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions And Answers On Thermochemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Questions And Answers On Thermochemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions And Answers On Thermochemistry
 - Personalized Recommendations
 - Multiple Choice Questions And Answers On Thermochemistry User Reviews and Ratings

- Multiple Choice Questions And Answers On Thermochemistry and Bestseller Lists
- 5. Accessing Multiple Choice Questions And Answers On Thermochemistry Free and Paid eBooks
 - Multiple Choice Questions And Answers On Thermochemistry Public Domain eBooks
 - Multiple Choice Questions And Answers On Thermochemistry eBook Subscription Services
 - Multiple Choice Questions And Answers On Thermochemistry Budget-Friendly Options
- 6. Navigating Multiple Choice Questions And Answers On Thermochemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Multiple Choice Questions And Answers On Thermochemistry Compatibility with Devices
 - Multiple Choice Questions And Answers On Thermochemistry Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Multiple Choice Questions And Answers On Thermochemistry
 - Highlighting and Note-Taking Multiple Choice Questions And Answers On Thermochemistry
 - Interactive Elements Multiple Choice Questions And Answers On Thermochemistry
- 8. Staying Engaged with Multiple Choice Questions And Answers On Thermochemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Questions And Answers On Thermochemistry
- 9. Balancing eBooks and Physical Books Multiple Choice Questions And Answers On Thermochemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Questions And Answers On Thermochemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Multiple Choice Questions And Answers On Thermochemistry
 - Setting Reading Goals Multiple Choice Questions And Answers On Thermochemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Multiple Choice Questions And Answers On Thermochemistry
 - Fact-Checking eBook Content of Multiple Choice Questions And Answers On Thermochemistry
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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