Sol. **2**

P – OH bonds = 4
P = O bonds = 2
P – O – P linkage = 1
Ans. 4, 2, 1

21. The mole fraction of glucose ($C_6H_{12}O_6$) in an aqueous binary solution is 0.1. The mass percentage of water in it, to the nearest integer, is _____.

Sol. **47 %**

$x_{Glucose} = 0.1$

$$\text{mass\% of glucose} = \frac{0.1 \times 180}{0.1 \times 180 + 0.9 \times 18} \times 100$$

$$= \frac{1800}{18 + 16.2}$$

$$= \frac{1800}{34.2} \%$$

$$= 52.63\%$$

$$= 53\%$$

$\therefore$ mass % of $H_2O = 47\%$

22. The volume strength of 8.9 M H_2O_2 solution calculated at 273 K and 1 atm is _____. ($R = 0.0821$ L atm K^{-1} mol^{-1}) (rounded off ot the nearest integer)

Sol. **100**

$$\text{Vol. strength} = \frac{8.9}{2} \times \frac{0.821 \times 273}{1}$$

$$= 99.73$$

$$= 100$$

23. An element with molar mass 2.7×10^{-2} kg mol^{-1} forms a cubic unit cell with edge length 405 pm. If its density is 2.7×10^3 kg m^{-3}, the radius of the element is approximately _____ $\times 10^{-12}$ m (to the nearest integer).

Sol. **143**

$$\text{Density} = \frac{Z \times GMM}{N_A \times a^3}$$

$$2.7 \times 10^3 = \frac{Z \times 2.7 \times 10^{-2}}{6.023 \times 10^{23} \times (405 \times 10^{-12})^3}$$

$Z = 6.023 \times 405 \times 405 \times 405 \times 10^{23-36+3+2}$

$Z = 6.023 \times 405 \times 405 \times 405 \times 10^{-8}$

$Z = 4$

FCC

$4R = \sqrt{2} \times a$

$R = \dfrac{405}{2\sqrt{2}} \times 10^{-12} = 143.21 \times 10^{-12}$m

$= 143$ ans

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Mamta Batra

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