

Chapter 5 Assessment

pages 166–169

Section 5.1

Mastering Concepts

- 34.** Define the following terms.
- frequency**
Frequency is the number of waves that pass a given point per second.
 - wavelength**
Wavelength is the shortest distance between equivalent points on a continuous wave.
 - quantum**
A quantum is the minimum amount of energy that can be lost or gained by an atom.
 - ground state**
An atom's ground state is its lowest allowable energy state.
- 35.** Arrange the following types of electromagnetic radiation in order of increasing wavelength.
- ultraviolet light
 - microwaves
 - radio waves
 - X rays
- d, X rays, a, ultraviolet light, b, microwaves, c, radio waves
- 36.** A gamma ray has a frequency of 2.88×10^{21} Hz. What does that mean?
- 2.88×10^{21} gamma-ray electromagnetic waves of this frequency pass a given point per second.
- 37.** What is the photoelectric effect?
- a phenomenon in which a metal emits electrons when light of a sufficient frequency shines on it
- 38. Neon Sign** How does light emitted from a neon sign differ from sunlight?
- The light from a neon sign contains only certain visible colors, while sunlight contains the full spectrum of colors.
- 39.** Explain Planck's quantum concept as it relates to energy lost or gained by matter.
- According to Planck, for a given frequency, ν , matter can emit or absorb energy only in discrete quanta that are whole-number multiples of $h\nu$, where h is Planck's constant.
- 40.** How did Einstein explain the photoelectric effect?
- He proposed that photons must have a certain minimum energy level, or threshold, value to cause the ejection of a photoelectron.
- 41. Rainbow** What are two differences between the red and green electromagnetic waves in a rainbow?
- The red waves have a longer wavelength and a lower frequency.
- 42. Temperature** What happens to the light emitted by a heated, glowing object as its temperature increases?
- The color of the light changes as the object acquires more and more energy.
- 43.** What are three deficiencies of the wave model of light related to light's interaction with matter?
- The wave model does not explain the photoelectric effect, atomic emission spectra, and why matter emits different frequencies of light at different temperatures.
- 44.** How are radio waves and ultraviolet waves similar? How are they different?
- Both types of waves travel at the same speed in a vacuum, 3.00×10^8 m/s. Radio waves have a longer wavelength and lower frequency than ultraviolet waves.

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Table of Contents Jchshub Solutions Manual

1. Understanding the eBook Jchshub Solutions Manual
 - The Rise of Digital Reading Jchshub Solutions Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Jchshub Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Jchshub Solutions Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Jchshub Solutions Manual
 - Personalized Recommendations
 - Jchshub Solutions Manual User Reviews and Ratings
 - Jchshub Solutions Manual and Bestseller Lists

5. Accessing Jchshub Solutions Manual Free and Paid eBooks
 - Jchshub Solutions Manual Public Domain eBooks
 - Jchshub Solutions Manual eBook Subscription Services
 - Jchshub Solutions Manual Budget-Friendly Options
6. Navigating Jchshub Solutions Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Jchshub Solutions Manual Compatibility with Devices
 - Jchshub Solutions Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Jchshub Solutions Manual
 - Highlighting and Note-Taking Jchshub Solutions Manual
 - Interactive Elements Jchshub Solutions Manual
8. Staying Engaged with Jchshub Solutions Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Jchshub Solutions Manual
9. Balancing eBooks and Physical Books Jchshub Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Jchshub Solutions Manual
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Jchshub Solutions Manual
 - Setting Reading Goals Jchshub Solutions Manual
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Jchshub Solutions Manual
 - Fact-Checking eBook Content of Jchshub Solutions Manual
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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