

1. Consider the following reaction:



An equilibrium mixture of this reaction at a certain temp was found to have  $[\text{CO}] = 0.45\text{M}$ ,  $[\text{H}_2] = 0.10\text{M}$ , and  $[\text{CH}_3\text{OH}] = 0.19\text{M}$ . What is the value of the equilibrium constant ( $K_c$ ) at this temperature?

$$K_c = \frac{[\text{CH}_3\text{OH}]}{[\text{CO}][\text{H}_2]^2} = \frac{(0.19)}{(0.45)(0.1)^2} = 136.9$$

2. Consider the following reaction:



An equilibrium mixture of this reaction at a certain temperature was found to have  $[\text{NH}_3] = 0.274\text{M}$  and  $[\text{H}_2\text{S}] = 0.355\text{M}$ . What is the value of the equilibrium constant ( $K_c$ ) at this temperature?

$$K_c = [\text{NH}_3][\text{H}_2\text{S}] = (0.274)(0.355) = 0.0973$$

3. Consider the following reaction:



The reaction between CO and  $\text{H}_2$  is carried out at a specific temperature with initial concentrations of  $[\text{CO}] = 0.32\text{M}$  and  $[\text{H}_2] = 0.45\text{M}$ . At equilibrium, the concentration of  $\text{CH}_3\text{OH}$  is  $0.14\text{M}$ . Find the equilibrium constant at this temperature.

	CO	$2\text{H}_2$	$\text{CH}_3\text{OH}$	$K_c = \frac{(0.14)}{(0.18)(0.17)^2} = 26.9$
I	0.32	0.45		
C	-x	-2x	+x	
E	0.18	0.17	0.14	

4. Consider the following reaction:



The reaction of  $\text{CH}_4$  is carried out at some temperature with an initial concentration of  $[\text{CH}_4] = 0.09\text{M}$ . At equilibrium, the concentration of  $\text{H}_2$  is  $0.012\text{M}$ . Find the equilibrium constant at this temperature.

	$2\text{CH}_4$	$\text{C}_2\text{H}_2$	$3\text{H}_2$	$K_c = \frac{(0.04)(0.012)^3}{(0.082)^2} = 1.0 \times 10^{-6}$
I	0.09			
C	-2x	+x	+3x	
E	0.082	0.004	0.012M	

5. Consider the following reaction:



A reaction mixture contains  $[\text{NO}_2] = 0.024\text{M}$  and  $[\text{N}_2\text{O}_4] = 0.033\text{M}$ . Calculate  $Q_c$  at the given concentrations.

$$Q_c = \frac{(0.024)^2}{(0.033)} = 0.0183 \quad Q_c > K_c, \text{ rxn goes left}$$

6. The following reaction was carried out in a 2.75L reaction vessel at  $1108\text{K}$ :



# Mastering Chemistry Answers Assignment 1

**Jean-Claude Falmagne, Dietrich  
Albert, Christopher Doble, David  
Eppstein, Xiangen Hu**



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## **Table of Contents Mastering Chemistry Answers Assignment 1**

1. Understanding the eBook Mastering Chemistry Answers Assignment 1
  - The Rise of Digital Reading Mastering Chemistry Answers Assignment 1
  - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Chemistry Answers Assignment 1
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mastering Chemistry Answers Assignment 1
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Chemistry Answers Assignment 1
  - Personalized Recommendations
  - Mastering Chemistry Answers Assignment 1 User Reviews and Ratings
  - Mastering Chemistry Answers Assignment 1 and Bestseller Lists
5. Accessing Mastering Chemistry Answers Assignment 1 Free and Paid eBooks
  - Mastering Chemistry Answers Assignment 1 Public Domain eBooks

- Mastering Chemistry Answers Assignment 1 eBook Subscription Services
- Mastering Chemistry Answers Assignment 1 Budget-Friendly Options
- 6. Navigating Mastering Chemistry Answers Assignment 1 eBook Formats
  - ePub, PDF, MOBI, and More
  - Mastering Chemistry Answers Assignment 1 Compatibility with Devices
  - Mastering Chemistry Answers Assignment 1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mastering Chemistry Answers Assignment 1
  - Highlighting and Note-Taking Mastering Chemistry Answers Assignment 1
  - Interactive Elements Mastering Chemistry Answers Assignment 1
- 8. Staying Engaged with Mastering Chemistry Answers Assignment 1
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mastering Chemistry Answers Assignment 1
- 9. Balancing eBooks and Physical Books Mastering Chemistry Answers Assignment 1
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mastering Chemistry Answers Assignment 1
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mastering Chemistry Answers Assignment 1
  - Setting Reading Goals Mastering Chemistry Answers Assignment 1
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mastering Chemistry Answers Assignment 1
  - Fact-Checking eBook Content of Mastering Chemistry Answers Assignment 1
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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by rearranging the equation we can see that the coefficient of  $x$  is the slope and the constant term is the  $y$  intercept in this case the coefficient of  $x$  is 2.3 so the slope is 2.3

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