

- (a) Anion is formed by the gain of electrons. Thus the numbers of electrons are more than protons. The effective positive charge in the nucleus is less, so less inward pull is experienced. Hence the size expands. So the size of an atom is greater than the size of parent atom.
- (b) Since Argon has stable octet configuration, so due to the inter- electronic repulsions the effect of nuclear pull over the valence shell electrons cannot be seen which results in the bigger size.
- (c) Since size of Bromine is bigger than chlorine, so it becomes more difficult for Br atoms to attract electrons. Thus, Cl is more reactive than Br.

Question 17:

Which element has:

- (a) two shells, both of which are completely filled with electrons?
- (b) the electronic configuration 2, 8, 3?
- (c) a total of three shells with five electrons in its valence shell?
- (d) a total of four shells with two electrons in its valence shell?
- (e) twice as many electrons in its second shell as in its first shell?

Solution 17:

- (a) Neon
- (b) Aluminum
- (c) Phosphorus
- (d) Calcium
- (e) Carbon

Question 18:

Name

- (a) An alkali metal in period 3 and halogen in period 2.
- (b) The noble gas with 3 shells.
- (c) The non-metals present in period 2 and metals in period 3.
- (d) The element of period 3 with valency 4
- (e) The element in period 3 which does not form oxide
- (f) The element of lower nuclear charge out of Be and Mg.
- (g) Which has higher E.A. fluorine or Neon.
- (h) Which has maximum metallic character Na, Li or K.

Solution 18:

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**Kaplan, Kaplan Educational Center
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