

CHEM 1201 - Mastering Chemistry HW 1 & 2 + Exam 1 Review (Ch. 1-3, Prof. Elgrishi) study guide verified A+ tips NEW!!!

1. **2, 1:** What is the ratio of hydrogen atoms (H) to oxygen atoms (O) in 2 LL of water? Enter the simplest whole number ratio in order of hydrogen to oxygen, respectively.
2. **600 H atoms:** How many atoms of hydrogen (H) are present in 200 molecules of ammonia (NH_3)?
3. **8.35×10^{25} H atoms:** A solution of ammonia and water contains 2.90×10^{25} water molecules and 8.50×10^{24} ammonia molecules. How many total hydrogen atoms are in this solution?
4. **Intensive:**
 - matte grey in color
 - begins to melt at 961.78 degrees Celcius
 - has a density of 10.49 g/mL**Extensive:**
 - has a volume of 7.42 mL
 - has a mass of 77.83 g**Classify the following as intensive or extensive properties of silver.**
 - matte grey in color
 - begins to melt at 961.78 degrees Celcius
 - has a density of 10.49 g/mL
 - has a volume of 7.42 mL
 - has a mass of 77.83 g
5. **the density of the sample**
the melting point of the sample: You are given a sample resembling silver. What properties could be used to help determine whether the sample is really silver?
6. **189 g:** By definition,
 $\text{density} = \text{mass} / \text{volume}$
Typical units of density are g/mL or g/cm^3 .
 $1 \text{ mL} = 1 \text{ cm}^3$
The density of iron is 7.87 g/mL. What is the mass of a $4.00 \text{ cm} \times 3.00 \text{ cm} \times 2.00 \text{ cm}$ block of iron?
7. **In moist air it combines with oxygen to form aluminum oxide.**

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M Walker



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