



$$\Delta H_{298}^\circ = -68.73 \text{ kcal } (-287.57 \text{ kJ})$$

Atomic Weight of **B** = 10.82 grams.

Atomic Weight of **O** = 16.0 grams.

Required:

- (a) Standard heat of formation ΔH_{298}° of B_2O_3 in terms of per mole of B_2O_3 =?
- (b) Standard heat of formation ΔH_{298}° of B_2O_3 in terms of per gram of B_2O_3 =?

Solution:

Now, we will multiply above given equations with real numbers to make them equal and then we will add all equations with each other. It will yield the final equation which will lead us to the final solution of this problem.

1. Multiply eq (i) with 2.



Metallurgical Thermodynamics Problems Solution

Victor M. Corman



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Electrochemical and Metallurgical Industry, 1921 **Fundamentals of Metallurgical Thermodynamics** Santosh Kumar Sahoo, Mithilesh Kumar, Swapan Kumar Karak, 2023-11-14 This book highlights introduction of thermodynamics first law second law third law of thermodynamics and their applications concepts of entropy free energies thermodynamic equilibrium thermodynamic activity and fugacity Maxwell relations Gibbs Helmholtz equation Clausius Clapeyron equation etc have been discussed in detail and made easily understandable to the undergraduate students of metallurgy Thermodynamics involved in formation of different types of solutions ideal real and regular solutions has also been discussed in detail This book also discusses the applications of various thermodynamic properties in different metallurgical operations At the end of each and every chapter different types of typical related problems have also been solved

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