

# Chapter 6

## Chapter 6: Thermochemistry

### 6.2 - Enthalpy and Calorimetry

Enthalpy

$$H = E + PV$$

For a chemical reaction, enthalpy change is given by the equation

$$\Delta H = H_{\text{products}} - H_{\text{reactants}}$$

Products of a reaction have a greater enthalpy ( $\Delta H$ ):

- $\Delta H$  is positive
- heat is absorbed by the system (endothermic reaction)

Reactants of a reaction have a greater enthalpy ( $\Delta H$ ):

- $\Delta H$  is negative
- heat is released by the system (exothermic reaction)

Calorimetry

Constant-pressure calorimetry

change in enthalpy = change in heat

$$\Delta H = q_p$$

# Mastering Chemistry Answers Chapter 6

**Norbert T. Porile**

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*Introductory Organic Biochemistry* Frederick A. Bettelheim,William M. Scovell,1997-08 *Introductory Chemistry* Charles H. Corwin,2011 To help students learn chemical skills and concepts more effectively *Introductory Chemistry Concepts and Critical Thinking* Sixth Edition highlights the connection between key concepts and key problem solving skills through critical thinking Math and problem solving are covered early in the text Corwin builds your problem solving ability through innovative learning aids and technology formulated to meet your needs This revision retains all the strengths of the previous editions while adding emphasis on conceptual understanding and critical thinking Instructor's Manual and Test

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