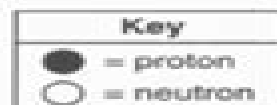


1. The diagram below represents the nucleus of an atom.



What are the atomic number and mass number of this atom?

- The atomic number is 9 and the mass number is 19.
  - The atomic number is 9 and the mass number is 20.
  - The atomic number is 11 and the mass number is 19.
  - The atomic number is 11 and the mass number is 20.
2. Which statement concerning elements is true?
- Different elements must have different numbers of isotopes.
  - Different elements must have different numbers of neutrons.
  - All atoms of a given element must have the same mass number.
  - All atoms of a given element must have the same atomic number.
3. Which subatomic particles are located in the nucleus of an He-4 atom?
- electrons and neutrons
  - electrons and protons
  - neutrons and protons
  - neutrons, protons, and electrons
4. Which part of a helium atom is positively charged?
- electron
  - neutron
  - nucleus
  - orbital
5. Which total mass is the *smallest*?
- the mass of 2 electrons
  - the mass of 2 neutrons
  - the mass of 1 electron plus the mass of 1 proton
  - the mass of 1 neutron plus the mass of 1 electron

- an uncharged nucleus
  - largely empty space
  - neutrons
  - protons
7. An experiment in which alpha particles were used to bombard thin sheets of gold foil led to the conclusion that an atom is composed mostly of
- empty space and has a small, negatively charged nucleus
  - empty space and has a small, positively charged nucleus
  - a large, dense, positively charged nucleus
  - a large, dense, negatively charged nucleus
8. Which conclusion is based on the "gold foil experiment" and the resulting model of the atom?
- An atom is mainly empty space, and the nucleus has a positive charge.
  - An atom is mainly empty space, and the nucleus has a negative charge.
  - An atom has hardly any empty space, and the nucleus has a positive charge.
  - An atom has hardly any empty space, and the nucleus has a negative charge.
9. Which statement about the mass of an electron is correct?
- The mass of an electron is equal to the mass of a proton.
  - The mass of an electron is less than the mass of a proton.
  - The mass of an electron is equal to the mass of a neutron.
  - The mass of an electron is greater than the mass of a neutron.
10. Experiments with gold foil indicated that atoms
- usually have a uniform distribution of positive charges
  - usually have a uniform distribution of negative charges
  - contain a positively charged, dense center
  - contain a negatively charged, dense center

# Mcq Pharm Chem Periodic Table

**Terry C. Jones**



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