

EQUILIBRIUM TEST

MULTIPLE CHOICE QUESTIONS

8) Consider the following reaction:



Which one of the following statements is correct?

- A. Since very large, the reaction is irreversible.
- B. Changing the pressure will affect the value of K_c .
- C. It is an very large, the reaction is effectively irreversible.
- D. Changing the temperature will not affect the value of K_c .

9) The following reaction, in which nitrogen trioxide, N_2O_3 , is converted to nitrogen dioxide, NO_2 , occurs at equilibrium:



In a given closed vessel, the equilibrium involved the three substances is 0.10 M. At this temperature, the initial concentration of N_2O_3 at equilibrium is 0.0500 M. At this temperature, the initial concentrations of NO_2 at equilibrium is

- A. 0.0500 M
- B. 1.000 M
- C. 1.40 M
- D. 0.00 M

10) Chloroacetic acid, ClCH_2COOH , is known with water, H_2O , as being weakly acidic. $\text{H}_2\text{ClCH}_2\text{COOH}$, $\text{H}_2\text{ClCH}_2\text{COO}^-$ equilibrium exist with H_2O as follows: $\text{H}_2\text{ClCH}_2\text{COOH}$ and hydrochloric acid, H_2Cl^+ according to the equilibrium below:



A solution that contains 0.1 M chloroacetic acid is placed in a vessel and left overnight. At that time the equilibrium exists. In the morning, after change is made in the system, which one of the following statements is most likely to result in an increase in the concentration of H_3O^+ ions?

- A. Change the solution.
- B. Bubble in more ClCH_2COOH gas.
- C. Add a few drops of phenolphthalein indicator.
- D. Add a few drops of 0.1 M sodium hydroxide, NaOH , solution.

Multiple Choice Concept Questions On Equilibrium

AW Rasmussen



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