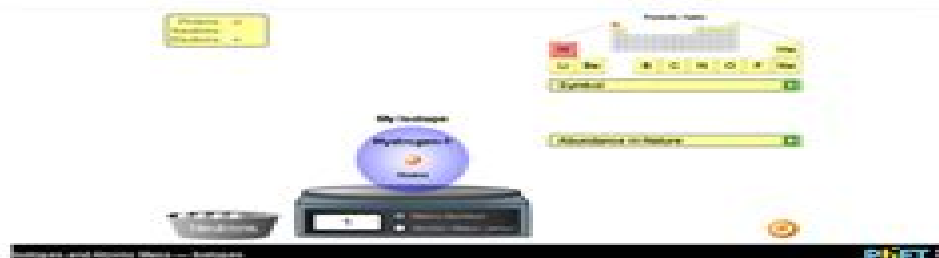


**Exploring Isotopes and Atomic Mass - [PhET Simulation](#) [ /60]**

**PART A - Making Isotopes:** Open the [Isotopes](#) screen, then explore the simulator for a few minutes to develop your own ideas about "isotopes".



1. What particles determine the mass number? [1]

The neutrons and protons are mainly the particles that determine the mass number and are found in the nucleus of the atom.

2. Why is mass number always a whole number? [1]

Mass number is always a whole number because it's the number of particles in the nucleus, there can't be fractions.

3. What is the approximate mass of one proton? [1]

The approximate mass of one proton is 1.00783 amu

4. What is the approximate mass of one neutron? [1]

The approximate mass of one neutron is 1.00627 amu

5. One isotope of carbon (C) has exactly the same mass number and atomic mass since it was used as the definition of the atomic mass unit (amu). Which isotope is it and what is its atomic mass? [2]

The isotope is Carbon-12 and its atomic mass is 12

6. Based on what you learned from the simulator, what is the definition of an isotope? *In your definition, include the terms: mass number, atomic number, number of protons ( $p^+$ ), neutrons ( $n^0$ ) and electrons ( $e^-$ ).* [3]

# Isotopes And Atomic Mass Phet Lab Answers

**J Dewey**



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