

Objective type

- Q.1 Which of the following is not a general characteristic of equilibria involving physical processes?
- (1) Equilibrium is possible only in a closed system at a given temperature.
 - (2) All measurable properties of the system remain constant.
 - (3) All the physical processes stop at equilibrium.
 - (4) The opposing processes occur at the same rate and there is dynamic but stable condition.
- Q.2 Which of the following statements is incorrect?
- (1) In equilibrium mixture of ice and water kept in perfectly insulated flask mass of ice and water does not change with time.
 - (2) The intensity of red colour increases when oxalic acid is added to a solution containing iron (III) nitrate and potassium thiocyanate.
 - (3) On addition of catalyst the equilibrium constant value is not affected.
 - (4) Equilibrium constant for a reaction with negative ΔH value decreases as the temperature increases.
- Q.3 When hydrochloric acid is added to cobalt nitrate solution at room temperature, the following reaction takes place and the reaction mixture becomes blue. On cooling the mixture it becomes pink. On the basis of this information mark the correct answer.
- $$\underset{\text{(pink)}}{[\text{Co}(\text{H}_2\text{O})_6]^{3+}}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \underset{\text{(blue)}}{[\text{CoCl}_4]^{2-}}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$$
- (1) $\Delta H > 0$ for the reaction
 - (2) $\Delta H < 0$ for the reaction
 - (3) $\Delta H = 0$ for the reaction
 - (4) The sign of ΔH cannot be predicted on the basis of this information.
- Q.4 What will be the correct order of vapour pressure of water, acetone and ether at 30°C . Given that among these compounds, water has maximum boiling point and ether has minimum boiling point?
- (1) Water < ether < acetone
 - (2) Water < acetone < ether
 - (3) Ether < acetone < water
 - (4) Acetone < ether < water
- Q.5 In which of the following reactions, the equilibrium remains unaffected on addition of small amount of argon at constant volume?
- (1) $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
 - (2) $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$
 - (3) $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
 - (4) The equilibrium will remain unaffected in all the three cases

Assertion & Reason

- Q.6 Assertion (A): In the dissociation of PCl_5 at constant pressure and temperature addition of helium at equilibrium increases the dissociation of PCl_5 .
- Reason (R) : Helium removes Cl_2 from the field of action.
- (1) Both A and R are true and R is correct explanation of A.
 - (2) Both A and R are true but R is not correct explanation of A.
 - (3) A is true but R is false.
 - (4) Both A and R are false.

Mcqs Of Environmental Chemistry

Gerard W. Vollmer



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Answers from the question itself If you are muddled by the given options then give each option a true and false test Instead of getting confused consider all the possibilities and neglect the incorrect options Hence in this way the most appropriate answer could be easily spotted Topic wise MCQs for NEET Chemistry Environmental Chemistry Organic Chemistry P Block Elements States of Matter Redox Reaction Chemical Equilibrium Thermodynamics Periodic Classification of Elements s Block Elements Chemical Bonding Molecular Structure Structure of Atom Some Basic Concepts of Chemistry Hydrocarbons Some Basic Concepts of Chemistry Solid State Structure of Atom Solutions Classification of Elements and Periodicity in Properties Electrochemistry Chemical Bonding and Molecular Structure Chemical Kinetics States of Matter Gases and Liquids Surface Chemistry Thermodynamics General Principles and Processes of Isolation of Elements Equilibrium Redox Reactions d and f Block Elements Hydrogen Coordination Compounds Haloalkanes and Haloarenes Some p Block Elements Alcohols Phenols and Ethers Organic Chemistry Some Basic Principles and Techniques Aldehyde Ketones and Carboxylic Acids Hydrocarbons Organic Compounds Containing Nitrogen Environmental Chemistry Biomolecules Polymers and Chemistry in Everyday Life Use a step wise approach to solve conceptual and complex questions Several times Matching type Questions are asked where the students are required to find the mismatched or the correctly matched option Some of the questions asked in the NEET exam are entirely memory based therefore the students are advised to memorize the common names scientific names concepts and important definitions Around 40% of the questions asked in the NEET exam are application based Therefore students need to focus more on the concepts along with its applications in order to score well in the examination Chemistry topics broadly can be grouped into Organic Inorganic and Physical chemistry Organic Chemistry is a vast section and holds the maximum weightage Its concepts and principles are important One needs to be conversant with properties of groups conversions recognizing organic compounds etc Interlinking between different chapters from organic chemistry can be of great help The students must primarily focus on reading NCERT textbooks Several times the questions asked in NEET exam are taken directly from the NCERT textbooks Initially avoid answering those questions for which you are not confident because your wrong answer may reduce your final score In order to utilize your time appropriately divide the three hours of examination time as per your comfort among Physics Chemistry and Biology Initially focus on attempting all easy questions and later on pick the difficult ones By this way your confidence will be elevated and you will also get more time to answer hard questions Practice previous years question papers mock tests and sample papers to get an idea on how to answer MCQ questions efficiently Preparing at an early stage is what an MCQ exam requires Avoid guesswork for negative marking questions as they might lower your final score These tips can be very helpful for students to answer difficult and brain teaser questions Prior preparations and practice are mandatory aspects of any examination Hence to crack highly competitive examination like NEET it is mandatory for students to prepare well and acquire the skills to tackle multiple choice questions effectively Rather than just following mere guesswork the aspirants can focus on the tips discussed to tackle Multiple Choice

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