

## CHEMICAL THERMODYNAMICS

- 1) In open system, system and surroundings exchange.....  
 (a) Energy only (b) Matter only  
**(c) Energy and matter** (d) Neither energy nor matter
- 2) An isothermal expansion takes place at.....  
 (a) constant pressure (b) constant volume  
**(c) constant temperature** (d) constant temperature & pressure
- 3) In a closed thermos flask few cubes of ice are placed. It is example of.....  
 (a) closed system (b) open system  
**(c) isolated system** (d) thermodynamics system
- 4) In an adiabatic compression, the internal energy of the system.....  
 (a) decreases **(b) increases** (c) increases or decreases (d) none
- 5) Identify the intensive quantities from the following.....  
 (a) volume (b) enthalpy **(c) pressure** (d) none
- 6) Which of the following is extensive property?  
 (a) mole (b) mass (c) volume **(d) all are correct**
- 7) The heat change in a chemical reaction at constant pressure is.....  
**(a)  $\Delta H$**  (b)  $\Delta U$  (c)  $\Delta T$  (d)  $\Delta V$
- 8)  $\Delta U$  is always positive when.....  
**(a) system absorbs heat and work is done on it**  
 (b) system emits heat and work is done on it  
 (c) system emits heat and no work is done on it  
 (d) system absorbs heat and work is done by it
- 9) If in a chemical reaction, the products have less energy than the reactants, that reaction is.....  
 (a) a reversible reaction **(b) an exothermic reaction**  
 (c) an isothermal reaction (d) an endothermic reaction
- 10) A small piece of Na metal is dropped into water. Na bursts into flame, the reaction is:  
 (a) neutralisation (b) reversible **(c) exothermic** (d) endothermic
- 11) For the reaction:  $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(l)$  at constant temperature,  $\Delta H - \Delta U$  is  
 (a)  $+RT$  **(b)  $-3RT$**  (c)  $+3RT$  (d)  $-RT$
- 12) When  $NH_4Cl$  is dissolved in water, the solution becomes cold. The change is?  
**(a) endothermic** (b) exothermic (c) super cooling (d) none

# Mcqs Of Thermodynamics With Answers

**SJ Ball**



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*Comprehensive Objective Physics* Narinder Kumar, Dr. J. K. Juneja, 2006

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**Problems In Chemical Thermodynamics, With Solutions** Maka Aleksishvili, Shota Sidamonidze, 2002-11-26 The methods of chemical thermodynamics are effectively used in many fields of science and technology Mastering these methods and their use in practice requires profound comprehension of the theoretical questions and acquisition of certain calculating skills This book is useful to undergraduate and graduate students in chemistry as well as chemical thermal and refrigerating technology it will also benefit specialists in all other fields who are interested in using these powerful methods in their practical activities

*Basic Thermodynamics* B. K. VENKANNA, SWATI B. V., B. K. Venkanna, B. V. Swati, 2010-07 This book provides an in depth discussion of the principles of thermodynamics It focuses on engineering applications of theory and sound techniques for solving thermodynamic problems The book presents the fundamental concepts of thermodynamics and describes the theory of work and heat The text covers in detail the first law and the second law of thermodynamics with their applications It also explains the concepts of entropy and availability and irreversibility In addition the book presents thermodynamic properties of pure substances ideal gases and mixtures of ideal gases as well as real gases This book is designed for undergraduate students of mechanical engineering industrial and production engineering automobile engineering and aeronautical engineering for their courses in thermodynamics Key Features Presents the text in a simple and elegant manner to enable the students to grasp the essentials of the subject easily and quickly Covers all types of problems of various difficulty levels Includes more than 300 worked out examples and a large number of end of chapter exercises Provides

solutions to several model question papers at the end of the book      *Thermodynamics* Rumi Michael Leigh, 2023-05-24

*Thermodynamics Things You Should Know Questions and Answers* explains the principles of thermodynamics in a question and answer format. The book introduces energy, work, heat, and the basic laws of thermodynamics. It describes the concepts of entropy, enthalpy, internal energy, and equilibrium, and shows how these ideas connect to physical and chemical processes. Key topics include the first and second laws of thermodynamics, the role of entropy in natural systems, and the behavior of gases as described by equations of state. The text also examines heat transfer, efficiency, and applications of thermodynamics in engines, refrigeration, and power systems. Worked examples and explanations show how thermodynamic concepts are applied in solving problems and understanding physical phenomena. This book will interest students, engineering learners, and those studying physics who want to understand thermodynamics and its applications.      Problems And Solutions On Thermodynamics And Statistical Mechanics Yung-kuo Lim, 1990-02-01

The material for these volumes has been selected from the past twenty years examination questions for graduate students at University of California at Berkeley, Columbia University, the University of Chicago, MIT, State University of New York at Buffalo, Princeton University, and University of Wisconsin.

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