

**1** Generally transition elements and their salts are coloured due to the presence of unpaired electrons in metal ions. Which of the following compounds are coloured?

- $\text{KMnO}_4$
- $\text{Ce}(\text{SO}_4)_2$
- $\text{TiCl}_3$
- $\text{Cu}_2\text{Cl}_2$

Choose the correct option

- (a, b)
- (b, c)
- (c, d)
- (a, d)

**2** General electronic configuration of actinoids is  $(n-2)f^{1-14}(n-1)d^{0-2}ns^2$ . Which of the following actinoids have one electron in 6d orbital?

- U (Atomic no. 92)
- Np (Atomic no. 93)
- Pu (Atomic no. 94)
- Am (Atomic no. 95)

Choose the correct option

- (a, b)
- (b, c)
- (c, d)
- (a, d)

**3** Which of the following lanthanoids show +2 oxidation state besides the characteristic oxidation state +3 of lanthanoids?

- Ce
- Eu
- Yb
- Ho

Choose the correct option

- (a, b)
- (b, c)
- (c, d)
- (a, d)

**4** On heating lead nitrate forms oxides of nitrogen and lead. The oxides formed are-

- $\text{N}_2\text{O}$ ,  $\text{PbO}$
- $\text{NO}_2$ ,  $\text{PbO}$
- $\text{NO}$ ,  $\text{PbO}$
- $\text{NO}$ ,  $\text{PbO}_2$

**5** A brown ring is formed in the ring test for  $\text{NO}_3^-$  ion. It is due to the formation of-

- $[\text{Fe}(\text{H}_2\text{O})_5(\text{NO})]^{2+}$
- $\text{FeSO}_4$ ,  $\text{NO}_2$
- $[\text{Fe}(\text{H}_2\text{O})_4(\text{NO})_2]^{2+}$
- $\text{FeSO}_4$ ,  $\text{HNO}_3$

**6** The oxidation state of the central atom in the anion of the compound  $\text{NaH}_2\text{PO}_2$  will be-

- +3
- +5
- +1
- 3

**7** A black compound of manganese reacts with a halogen acid to give greenish-yellow gas. When an excess of this gas reacts with  $\text{NH}_3$  an unstable trihalide is formed. In this process, the oxidation state of nitrogen changes from-

- 3 to +3
- 3 to 0
- 3 to +5
- 0 to -3

**8** Which of the following is isoelectronic pair?

- $\text{ICl}_2$ ,  $\text{ClO}_2$
- $\text{BrO}_2^-$ ,  $\text{BrF}_2^+$
- $\text{ClO}_2$ ,  $\text{BrF}$
- $\text{CN}^-$ ,  $\text{O}_3$

**9** Which of the following statements are correct for  $\text{SO}_2$  gas?

- It acts as a bleaching agent in moist conditions.
  - Its molecule has linear geometry.
  - Its dilute solution is used as a disinfectant.
  - It can be prepared by the reaction of dilute  $\text{H}_2\text{SO}_4$  with metal sulphide.
- (a, b)
  - (b, c)
  - (c, d)
  - (a, c)

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