

Acids and Bases

Acid				Base			
Strength of Acid	K _a	Name	Formula	Formula	Name	Strength of Base	
Strongest						Weakest	
	Large	Perchloric acid	HClO ₄	ClO ₄ ⁻	Perchlorate ion		
	3.2×10 ⁸	Hydroiodic acid		I ⁻	Iodide		
	1.0×10 ⁹	Hydrobromic acid	HBr	Br ⁻	Bromide		
	1.3×10 ⁴	Hydrochloric acid	HCl	Cl ⁻	Chloride		
	1.0×10 ³	Sulfuric acid	H ₂ SO ₄	HSO ₄ ⁻	Hydrogen sulfate ion		
	2.4×10 ¹	Nitric acid	HNO ₃	NO ₃ ⁻	Nitrate ion		
	-----	Hydronium ion	H ₃ O ⁺	H ₂ O	Water		
	5.4×10 ⁻²	Oxalic acid	HO ₂ C ₂ O ₂ H	HO ₂ C ₂ O ₂ ⁻	Hydrogen oxalate ion		
	1.3×10 ⁻²	Sulfurous acid	H ₂ SO ₃	HSO ₃ ⁻	Hydrogen sulfite ion		
	1.0×10 ⁻²	Hydrogen sulfate ion	HSO ₄ ⁻	SO ₄ ²⁻	Sulfate ion		
	7.1×10 ⁻³	Phosphoric acid	H ₃ PO ₄	H ₂ PO ₄ ⁻	Dihydrogen phosphate ion		
	7.2×10 ⁻⁴	Nitrous acid	HNO ₂	NO ₂ ⁻	Nitrite ion		
	6.6×10 ⁻⁴	Hydrofluoric acid	HF	F ⁻	Fluoride ion		
	1.8×10 ⁻⁴	Methanoic acid	HCO ₂ H	HCO ₂ ⁻	Methanoate ion		
	6.3×10 ⁻⁵	Benzoic acid	C ₆ H ₅ COOH	C ₆ H ₅ COO ⁻	Benzoate ion		
	5.4×10 ⁻⁵	Hydrogen oxalate ion	HO ₂ C ₂ O ₂ ⁻	O ₂ C ₂ O ₂ ²⁻	Oxalate ion		
	1.8×10 ⁻⁵	Ethanoic acid	CH ₃ COOH	CH ₃ COO ⁻	Ethanoate (acetate) ion		
	4.4×10 ⁻⁷	Carbonic acid	H ₂ CO ₃	HCO ₃ ⁻	Hydrogen carbonate ion		
	1.1×10 ⁻⁷	Hydrosulfuric acid	H ₂ S	HS ⁻	Hydrogen sulfide ion		
	6.3×10 ⁻⁸	Dihydrogen phosphate ion	H ₂ PO ₄ ⁻	HPO ₄ ²⁻	Hydrogen phosphate ion		
	6.2×10 ⁻⁸	Hydrogen sulfite ion	HS ⁻	SO ₃ ²⁻	Sulfite ion		
	2.9×10 ⁻⁸	Hypochlorous acid	HClO	ClO ⁻	Hypochlorite ion		
	6.2×10 ⁻¹⁰	Hydrocyanic acid	HCN	CN ⁻	Cyanide ion		
	5.8×10 ⁻¹⁰	Ammonium ion	NH ₄ ⁺	NH ₃	Ammonia		
	5.8×10 ⁻¹⁰	Boric acid	H ₃ BO ₃	H ₂ BO ₃ ²⁻	Dihydrogen borate		
	4.7×10 ⁻¹¹	Hydrogen carbonate ion	HCO ₃ ⁻	CO ₃ ²⁻	Carbonate ion		
	4.2×10 ⁻¹³	Hydrogen phosphate ion	HPO ₄ ²⁻	PO ₄ ³⁻	Phosphate ion		
	1.8×10 ⁻¹³	Dihydrogen borate ion	H ₂ BO ₃ ⁻	HBO ₃ ²⁻	Hydrogen borate ion		
	1.3×10 ⁻¹³	Hydrogen sulfide ion	HS ⁻	S ²⁻	Sulfide ion		
	1.6×10 ⁻¹⁴	Hydrogen borate ion	HBO ₃ ²⁻	BO ₃ ²⁻	Borate ion		
Weakest	-----	water	H ₂ O	OH ⁻	Hydroxide	Strongest	

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