

Limiting Reactants Practice KEY

1. How many grams of aluminum chloride could be produced from 35.5 g of aluminum and 39.0 g of chlorine gas?



$$35.5 \text{ g Al} \times \frac{1 \text{ mol Al}}{26.982 \text{ g Al}} \times \frac{3 \text{ mol Cl}_2}{2 \text{ mol Al}} \times \frac{70.906 \text{ g Cl}_2}{1 \text{ mol Cl}_2} = 140. \text{ g Cl}_2$$

- ∞ This means 35.5 g Al would react with 140. g Cl₂
- ∞ Only 39.0 g Cl₂ were available, so it is the limiting reactant.
- ∞ Use the mass of Cl₂ in the reaction to determine the mass of the AlCl₃ formed.

$$39.0 \text{ g Cl}_2 = \text{ ? } \text{ g AlCl}_3$$

$$39.0 \text{ g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.906 \text{ g Cl}_2} \times \frac{2 \text{ mol AlCl}_3}{3 \text{ mol Cl}_2} \times \frac{133.341 \text{ g AlCl}_3}{1 \text{ mol AlCl}_3} =$$

Answer: 48.9 g AlCl₃

2. Identify the limiting reagent when 10.8 g water reacts with 4.7 g Na to produce NaOH and H₂.

To solve, set one reactant to the other and compare what is calculated to the actual amount in the reaction to determine the theoretical limiting reactant.

$$10.8 \text{ g H}_2\text{O} = \text{ ? } \text{ g Na}$$

Or

$$4.7 \text{ g Na} = \text{ ? } \text{ g H}_2\text{O}$$



$$10.8 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.015 \text{ g H}_2\text{O}} \times \frac{2 \text{ mol Na}}{2 \text{ mol H}_2\text{O}} \times \frac{22.990 \text{ g Na}}{1 \text{ mol Na}} = 13.8 \text{ g Na}$$

- ∞ This means that 10.8 g will react completely with 13.8 g Na, but only 4.7 g Na was available to this reaction. Therefore, Na is the limiting reactant in this problem.

Limiting Reactant Problems Answer Key

Darrell D. Ebbing



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Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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