

## Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)



a) If you start with 14.8 g of  $\text{C}_3\text{H}_8$  and 3.44 g of  $\text{O}_2$ , determine the limiting reagent

$$14.8 \text{ g C}_3\text{H}_8 \times \frac{1 \text{ mol C}_3\text{H}_8}{44.09 \text{ g C}_3\text{H}_8} \times \frac{4 \text{ mol H}_2\text{O}}{1 \text{ mol C}_3\text{H}_8} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.2 \text{ g H}_2\text{O}$$

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{4 \text{ mol H}_2\text{O}}{5 \text{ mol O}_2} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 1.55 \text{ g H}_2\text{O}$$

$\text{O}_2$  is limiting

b) determine the number of moles of carbon dioxide produced

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{3 \text{ mol CO}_2}{5 \text{ mol O}_2} \times \frac{44.01 \text{ g CO}_2}{1 \text{ mol CO}_2} = 2.84 \text{ g H}_2\text{O}$$

c) determine the number of grams of  $\text{H}_2\text{O}$  produced

$$1.55 \text{ g H}_2\text{O}$$

d) determine the number of grams of excess reagent left

$$3.44 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{32 \text{ g O}_2} \times \frac{1 \text{ mol C}_3\text{H}_8}{5 \text{ mol O}_2} \times \frac{44.09 \text{ g C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} = .948 \text{ g C}_3\text{H}_8 \text{ used}$$

$$14.8 \text{ g C}_3\text{H}_8 - 0.948 \text{ g C}_3\text{H}_8 = 13.9 \text{ g C}_3\text{H}_8 \text{ remain}$$

2. Given the following equation:



a) If 10.0 g of  $\text{Al}_2(\text{SO}_4)_3$  is reacted with 10.0 g of  $\text{NaOH}$ , determine the limiting reagent

$$10.00 \text{ g Al}_2(\text{SO}_4)_3 \times \frac{1 \text{ mol Al}_2(\text{SO}_4)_3}{294.17 \text{ g Al}_2(\text{SO}_4)_3} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 12.9 \text{ g Na}_2\text{SO}_4$$

$$10.00 \text{ g NaOH} \times \frac{1 \text{ mol NaOH}}{294.17 \text{ g NaOH}} \times \frac{3 \text{ mol Na}_2\text{SO}_4}{6 \text{ mol NaOH}} \times \frac{126.05 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} = 15.8 \text{ g Na}_2\text{SO}_4$$

# Limiting Reagents Answer Key Chemistry

**Stephen DeMeo**



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Chemistry John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind. A reorganized Table of Contents and inclusion of SI units, IUPAC standards, and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings. It more accurately reflects the curriculum of most Canadian institutions. Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant.

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*General Chemistry* Raymond Chang, 1986 *College Chemistry* Henry Fuller Holtzclaw, William R. Robinson, William Harrison Nebergall, 1984

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