

2. Limiting and Excess Reactants

The **limiting reactant** is defined as the species in a chemical reaction that **theoretically** would be the first to be completely consumed if the reaction were to proceed to completion according to the chemical equation.

- even if the reaction does not proceed to completion!

All of the other reactant/s are called **excess reactants**.

For Example:



If 1 g mol of C_7H_{16} and 12 g mol of O_2 are mixed so as to react,

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RD Boyd



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