

Name : \_\_\_\_\_

Score : \_\_\_\_\_ Date : \_\_\_\_\_

## Segment Lengths in Circles

Determine the length of segment(s) in the following circles

1



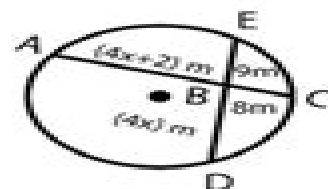
AB = \_\_\_\_\_

2



XY = \_\_\_\_\_

3



AB = \_\_\_\_\_, BD = \_\_\_\_\_

4



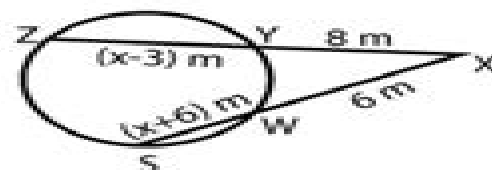
ST = \_\_\_\_\_

5



AB = \_\_\_\_\_

6



ZY = \_\_\_\_\_, SW = \_\_\_\_\_

7



RG = \_\_\_\_\_

8



GH = \_\_\_\_\_

# Lines That Intersect Circles Answer Key

**Daniel F McAuley**



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web solution average atomic mass of li 6 941 amu x 6 0151 amu 1 x 7 0160 amu 6 941 1 0009 x 7 0160 1 0009 x 0 075 x 0 075 x 0 075 corresponds to a natural abundance of 6li of 7 5 percent the natural abundance of 7li is x 0 925 or 92 5 percent 3 7 the unit factor required is 6 022

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web a molar mass  $Li_2CO_3$  2 6 941 g 12 01 g 3 16 00 g 73 89 g b molar mass  $CS_2$  12 01 g 2 32 07 g 76 15 g c molar mass  $CHCl_3$  12 01 g 1 008 g 3 35 45 g 119 37 g d molar mass  $C_6H_8O_6$  6 12 01 g 8 1 008 g 6 16 00 g 176 12 g e molar mass  $KNO_3$  39 10 g 14 01 g 3 16 00 g 101 11 g

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web it is a small step from mole mass calculations to mass mass calculations if we start with a known mass of one substance in a chemical reaction instead of a known number of moles we can calculate the corresponding masses of other substances in the reaction

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web jul 4 2022 to calculate the mass of titanium metal that you can obtain multiply the number of moles of titanium by the molar mass of titanium  $47.867 \text{ g/mol}$  moles ti mass ti times molar mass ti  $4.12 \text{ mol ti} \times 47.867 \text{ g/mol ti} = 197 \text{ g ti}$  notag

chapter 3 mass relationships in chemical reactions - Nov 23 2022

web 41 42 chemistry ch 3 mass relationships in chemical reactions molecular mass of  $\text{NO}_2$  atomic mass of n 2 atomic mass of o 14 01 ainu 2 16 00 ainu 46 01 amu atomic mass of  $\text{N}_2\text{O}_5$  2 atomic mass of n 5 atomic mass of o 2 14 01 ainu 5 16 00 ainu 108 02 amu example 3 2 exercise 3 3