

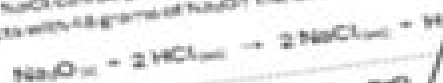
Stoichiometry

● Limiting Reactant & Percentage Yield ●

Answers

Stoichiometry (Page 8) Level 5: Limiting Reactants

- 5) How many grams of NaCl can be produced from the reaction when 50 g solution of HCl reacts with 48 grams of NaOH? The chemical equation is



Mole of NaOH	Mole	NaOH	HCl	NaCl
$n = \frac{m}{M}$	$n = 0.775 \text{ mol}$			
		48 g NaOH used up	5.275	1.54
Mole of HCl	$n = 0.975 \text{ mol}$	50 g HCl used up	5.425	5.85

- 6) How many liters of CO_2 at STP can be produced when 10 g benzene (C_6H_6) reacts with 52 g O_2 ? The chemical equation is unbalanced.



Mole of C_6H_6	Mole	C_6H_6	O_2
$n = \frac{m}{M}$	$n = 0.0775 \text{ mol}$	10 g C_6H_6	
Mole of O_2	$n = 1.625 \text{ mol}$	52 g O_2	1.625

Answers

Stoichiometry (Page 12) Level 6: Percentage Yield

1. If 15.8 grams of methane (CH_4) undergo complete combustion in excess oxygen to produce 29.4 grams of water, what is the percentage yield of this reaction? The chemical equation is balanced.



$n = \frac{m}{M}$	$n = 0.9875 \text{ mol}$	$\text{CH}_4 + \text{H}_2\text{O}$	$n = \frac{m}{M}$
$n = \frac{15.8 \text{ g}}{16.04 \text{ g/mol}}$		$1 \rightarrow 2$	$1.97 = \frac{m}{18.02}$
$n = 0.9875 \text{ mol}$		$0.9875 \rightarrow x$	$m = 35.5 \text{ grams}$
	$n = 1.97 \text{ mol}$		

Percentage (%) Yield

$\% = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$

Percentage (%) Yield

$$\% = \frac{\text{Actual}}{\text{Theoretical}} \times 100 \quad \% = \frac{29.4}{26.3} \times 100 = 111.8\%$$

Answer: **111.8%**

2. If 78.3 grams of sodium carbonate (Na_2CO_3) reacts with 822 mL of 1.12 M solution of hydrochloric acid to produce 49.7 grams of sodium chloride, what is the limiting reactant and the percentage yield of this reaction? The chemical equation is unbalanced.



Mole of Na_2CO_3	Mole	Na_2CO_3	HCl	NaCl	CO_2	H_2O
$n = \frac{m}{M}$	$n = 0.375 \text{ mol}$	78.3 g Na_2CO_3				
Mole of HCl	$n = 0.926 \text{ mol}$		822 mL	9.26 g		

Answer Key available

Limiting Reactant: **Hydrochloric acid**

Limiting Reactants And Percentage Yield Quiz

Purshotam Kaushik & Kirti Kaushik



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