

Limiting & Excess Reactants

Cheney, G., Thomas, T., McClelland, M., Salovey, M., Sluyter, P., & Mather, J. (2003). *Emotion: A handbook of research and applications*. Mahwah, NJ: Lawrence Erlbaum.

1. *Journal of the American Medical Association*, 2000; 284: 2689-2695.

- 1.1. **Redox reactions:** $\text{Fe}^{2+} + \text{H}_2\text{O}_2 + 2\text{H}^+ \rightarrow \text{Fe}^{3+} + 2\text{H}_2\text{O}$
In aqueous solution, Fe^{2+} is oxidized to Fe^{3+} by H_2O_2 .
Overall equation:
$$\text{Fe}^{2+} + \text{H}_2\text{O}_2 + 2\text{H}^+ \rightarrow \text{Fe}^{3+} + 2\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{H}_2\text{O}_2 + 2\text{H}^+ + 2e^- \rightarrow 2\text{H}_2\text{O}$
- 1.2. **Reaction of Fe^{2+} with KMnO_4 in acidic solution:**
 $\text{Fe}^{2+} + \text{KMnO}_4 + \text{H}^+ \rightarrow \text{Fe}^{3+} + \text{Mn}^{2+} + \text{H}_2\text{O}$
In acidic solution, Fe^{2+} is oxidized to Fe^{3+} and KMnO_4 is reduced to Mn^{2+} .
Overall equation:
$$5\text{Fe}^{2+} + \text{MnO}_4^- + 8\text{H}^+ \rightarrow 5\text{Fe}^{3+} + \text{Mn}^{2+} + 4\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{MnO}_4^- + 8\text{H}^+ + 5e^- \rightarrow \text{Mn}^{2+} + 4\text{H}_2\text{O}$
- 1.3. **Reaction of Fe^{2+} with $\text{Cr}_2\text{O}_7^{2-}$ in acidic solution:**
 $\text{Fe}^{2+} + \text{Cr}_2\text{O}_7^{2-} + \text{H}^+ \rightarrow \text{Fe}^{3+} + \text{Cr}^{3+} + \text{H}_2\text{O}$
In acidic solution, Fe^{2+} is oxidized to Fe^{3+} and $\text{Cr}_2\text{O}_7^{2-}$ is reduced to Cr^{3+} .
Overall equation:
$$6\text{Fe}^{2+} + \text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ \rightarrow 6\text{Fe}^{3+} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{Cr}_2\text{O}_7^{2-} + 14\text{H}^+ + 6e^- \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$
- 1.4. **Reaction of Fe^{2+} with IO_3^- in acidic solution:**
 $\text{Fe}^{2+} + \text{IO}_3^- + \text{H}^+ \rightarrow \text{Fe}^{3+} + \text{I}_2 + \text{H}_2\text{O}$
In acidic solution, Fe^{2+} is oxidized to Fe^{3+} and IO_3^- is reduced to I_2 .
Overall equation:
$$10\text{Fe}^{2+} + 2\text{IO}_3^- + 12\text{H}^+ \rightarrow 10\text{Fe}^{3+} + \text{I}_2 + 6\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $2\text{IO}_3^- + 12\text{H}^+ + 10e^- \rightarrow \text{I}_2 + 6\text{H}_2\text{O}$
- 1.5. **Reaction of Fe^{2+} with MnO_4^- in basic solution:**
 $\text{Fe}^{2+} + \text{MnO}_4^- + \text{OH}^- \rightarrow \text{Fe}^{3+} + \text{MnO}_2 + \text{H}_2\text{O}$
In basic solution, Fe^{2+} is oxidized to Fe^{3+} and MnO_4^- is reduced to MnO_2 .
Overall equation:
$$3\text{Fe}^{2+} + \text{MnO}_4^- + 4\text{OH}^- \rightarrow 3\text{Fe}^{3+} + \text{MnO}_2 + 2\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{MnO}_4^- + 2\text{H}_2\text{O} + 3e^- \rightarrow \text{MnO}_2 + 4\text{OH}^-$
- 1.6. **Reaction of Fe^{2+} with $\text{Cr}_2\text{O}_7^{2-}$ in basic solution:**
 $\text{Fe}^{2+} + \text{Cr}_2\text{O}_7^{2-} + \text{OH}^- \rightarrow \text{Fe}^{3+} + \text{Cr}^{3+} + \text{H}_2\text{O}$
In basic solution, Fe^{2+} is oxidized to Fe^{3+} and $\text{Cr}_2\text{O}_7^{2-}$ is reduced to Cr^{3+} .
Overall equation:
$$6\text{Fe}^{2+} + \text{Cr}_2\text{O}_7^{2-} + 14\text{OH}^- \rightarrow 6\text{Fe}^{3+} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{Cr}_2\text{O}_7^{2-} + 14\text{OH}^- + 6e^- \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$
- 1.7. **Reaction of Fe^{2+} with IO_3^- in basic solution:**
 $\text{Fe}^{2+} + \text{IO}_3^- + \text{OH}^- \rightarrow \text{Fe}^{3+} + \text{I}_2 + \text{H}_2\text{O}$
In basic solution, Fe^{2+} is oxidized to Fe^{3+} and IO_3^- is reduced to I_2 .
Overall equation:
$$10\text{Fe}^{2+} + 2\text{IO}_3^- + 12\text{OH}^- \rightarrow 10\text{Fe}^{3+} + \text{I}_2 + 6\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $2\text{IO}_3^- + 12\text{OH}^- + 10e^- \rightarrow \text{I}_2 + 6\text{H}_2\text{O}$
- 1.8. **Reaction of Fe^{2+} with MnO_4^- in basic solution:**
 $\text{Fe}^{2+} + \text{MnO}_4^- + \text{OH}^- \rightarrow \text{Fe}^{3+} + \text{MnO}_2 + \text{H}_2\text{O}$
In basic solution, Fe^{2+} is oxidized to Fe^{3+} and MnO_4^- is reduced to MnO_2 .
Overall equation:
$$3\text{Fe}^{2+} + \text{MnO}_4^- + 4\text{OH}^- \rightarrow 3\text{Fe}^{3+} + \text{MnO}_2 + 2\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{MnO}_4^- + 2\text{H}_2\text{O} + 3e^- \rightarrow \text{MnO}_2 + 4\text{OH}^-$
- 1.9. **Reaction of Fe^{2+} with $\text{Cr}_2\text{O}_7^{2-}$ in basic solution:**
 $\text{Fe}^{2+} + \text{Cr}_2\text{O}_7^{2-} + \text{OH}^- \rightarrow \text{Fe}^{3+} + \text{Cr}^{3+} + \text{H}_2\text{O}$
In basic solution, Fe^{2+} is oxidized to Fe^{3+} and $\text{Cr}_2\text{O}_7^{2-}$ is reduced to Cr^{3+} .
Overall equation:
$$6\text{Fe}^{2+} + \text{Cr}_2\text{O}_7^{2-} + 14\text{OH}^- \rightarrow 6\text{Fe}^{3+} + 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $\text{Cr}_2\text{O}_7^{2-} + 14\text{OH}^- + 6e^- \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$
- 1.10. **Reaction of Fe^{2+} with IO_3^- in basic solution:**
 $\text{Fe}^{2+} + \text{IO}_3^- + \text{OH}^- \rightarrow \text{Fe}^{3+} + \text{I}_2 + \text{H}_2\text{O}$
In basic solution, Fe^{2+} is oxidized to Fe^{3+} and IO_3^- is reduced to I_2 .
Overall equation:
$$10\text{Fe}^{2+} + 2\text{IO}_3^- + 12\text{OH}^- \rightarrow 10\text{Fe}^{3+} + \text{I}_2 + 6\text{H}_2\text{O}$$

Half-reactions:
Oxidation: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + e^-$
Reduction: $2\text{IO}_3^- + 12\text{OH}^- + 10e^- \rightarrow \text{I}_2 + 6\text{H}_2\text{O}$

Limiting And Excess Reactants Packet Model 1

Rachel S Tattersall



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