

**Limiting Reagent & Percent Yield Practice Worksheet**

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



- b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

**Only 13 g can be formed**

- c. What is the limiting reagent for the reaction in #2? copper (II) chloride

- d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu(NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu(NO}_3)_2}{1 \text{ mol Cu(NO}_3)_2} = 20.9 \text{ g Cu(NO}_3)_2$$

- e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol CuCl}_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

- f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



- b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.214 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb(NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

**8.51 g NaNO<sub>3</sub> can be made**

# Limiting Reagent And Percent Yield Answer Key

**Marian DeWane, Heather Hattori**



## **Limiting Reagent And Percent Yield Answer Key:**

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chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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**Chemical Principles** William L. Masterton, Emil J. Slowinski, Conrad L. Stanitski, 1981 **Holt Chemistry File**, 1998 This reference is a must for students who need extra help reteaching or extra practice The guide moves students through the same concepts as the text but at a slower pace More descriptive detail along with visual algorithms provides a more structured approach Each chapter closes with a large bank of practice problems Book jacket **Chemistry** McGraw-Hill Staff, 2001-03 *Practice Makes Perfect Chemistry Review and Workbook, Second Edition* Marian DeWane, Heather Hattori, 2018-12-28 The Winning Equation for Success in Chemistry is Practice Practice Practice This book will help you apply concepts and see how chemistry topics are interconnected Inside are numerous lessons to help you better understand the subject These lessons are accompanied by dozens of exercises to practice what you've learned along with a complete answer key to check your work Throughout this book you will learn the terms to help you understand chemistry and you will expand your knowledge of the subject through hundreds of sample questions and their solutions With the lessons in this book you will find it easier than ever to grasp chemistry concepts And with a variety of exercises for practice you will gain confidence using your growing chemistry skills in your classwork and on exams YOU LL BE ON YOUR WAY TO MASTERING THESE TOPICS AND MORE Atomic structure The periodic table Chemical formulas Chemical reactions Mass and mole relationships Gas laws Solutions Acids and bases Thermochemistry A brand new chapter on the structure of molecules Chemistry Success in 20 Minutes a Day Michael B. McGinnis, 2005 Offers a diagnostic test and twenty lessons covering vital chemistry skills Introductory Chemistry for Today Spencer L. Seager, Michael R. Slabaugh, 2004 Distinguished by its superior allied health focus and integration of technology Seager and Slabaugh's INTRODUCTORY CHEMISTRY FOR TODAY Fifth Edition continues to lead the market on both fronts through numerous allied health related applications examples boxes and a new Companion Web Site GOB ChemistryNow™ In addition to the many resources found in GOB ChemistryNow this powerful new Web site contains questions modeled after the Nursing School and Allied Health Entrance Exams and NCLEX LPN Certification Exams The authors strive to dispel users inherent fear of chemistry and to instill an appreciation for the role chemistry plays in our daily lives through a rich pedagogical structure and an accessible writing style that provides lucid explanations In addition Seager and Slabaugh's CHEMISTRY FOR TODAY Fifth Edition provides greater support in both problem solving and critical thinking skills By demonstrating how this information will be important to a reader's future career and providing important career information online the authors not only help readers to set goals but also to focus on

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1 the safest way to remove chips from a lathe or milling machine brush 2 the most common injury from using a hand file being stabbed by the tang file without a handle 3 msds stands for material safety data sheet see page 707 4 hazardous material labeling see page 53 57 a nfpa stands for national fire protection association b

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free groups branches or sections can be established which nims management characteristic does this describe a modular organization b management by objectives c comprehensive resource management d