

Name: _____

Date: _____

Student Exploration: Limiting Reactants

Go to www.explorelearning.com. Click "login with student code" and type the code ALEH8807P2.

Vocabulary: chemical equation, chemical formula, chemical reaction, coefficient, limiting reactant, molecule, product, reactant, subscript.

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

Imagine you and your friends are making hot dogs. A complete hot dog consists of a roller and a bun. At the store, you buy four packages of eight rollers and three bags of 10 buns.

1. How many total hot dogs can you make? _____
2. Which ingredient limited the number of hot dogs you could make? _____
3. Which ingredient will you have leftovers of? _____

Gizmo Warm-up

Just as ingredients can be put together to make a new food, substances can combine during a **chemical reaction** to produce new substances. The substances that undergo change are called **reactants**. The new substances are **products**.

Sometimes during a chemical reaction, one type of reactant will be used up before the other reactants. This reactant is the **limiting reactant**. Using the Limiting Reactants Gizmo™, you can determine which reactant is limiting in various scenarios.



To begin, make sure $\text{H}_2 + \text{O}_2$ becomes H_2O is selected. The small "2" in H_2 , O_2 , and H_2O is a **subscript**. Subscripts represent the number of atoms in a **molecule**.

1. Use the sliders to set the number of O_2 molecules and H_2 molecules to two.
 - A. How many hydrogen molecules (H_2) are there? _____ Hydrogen atoms? _____
 - B. How many oxygen molecules (O_2) are there? _____ Oxygen atoms? _____
2. How many H_2O molecules do you think will form when these four molecules react? _____
3. Click **Play** (▶). How many H_2O molecules actually formed? _____

Limiting Reactants Answer Key

Howard Stephen Stoker



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Foundations of Physical Science Tom Hsu PH.D., 2002

General Chemistry Key Concise Review Guide Sixth Edition Darrell D. Ebbing, 1999

Holt Chemistry Ralph Thomas Myers, 2004

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Introduction to Chemical Principles Howard Stephen Stoker, 1990 Introduction to Chemical Principles is a text for students who have had little to no previous instruction in chemistry or who had such instruction long enough ago that a thorough review is needed preface

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Preparatory Chemistry Howard Stephen Stoker, 1990

Student Solutions Manual: Ssm Chemistry Deborah Wiegand, 2003 This manual contains complete worked out solutions to all follow up problems and about half of all the chapter problems Each chapter of solutions opens with a summary of the text chapter content and a list of key equations needed to solve the problems

Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook McGraw Hill, 2001-08 Glencoe Chemistry Solving Problems A

Chemistry Handbook Matter and Change *SourceBook Version 2.1* ,1998 **Prentice Hall Chemistry** ,2000

General Chemistry Ralph H. Petrucci, William S. Harwood, 1993 General Chemistry Principles and Modern Applications is recognized for its superior problems lucid writing and precision of argument This updated and expanded edition retains the popular and innovative features of previous editions including Feature Problems follow up Integrative and Practice Exercises to accompany every in chapter Example and Focus On application boxes as well as new Keep in Mind marginal notes Topics covered include atoms and the atomic theory chemical compounds and reactions gases Thermochemistry electrons in atoms chemical bonding liquids solids and intermolecular forces chemical kinetics principles of chemical equilibrium acids and bases electrochemistry representative and transitional elements and nuclear and organic chemistry For individuals interested in a broad overview of chemical principles and applications **Must Know High School**

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section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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Table of Contents Limiting Reactants Answer Key

1. Understanding the eBook Limiting Reactants Answer Key
 - The Rise of Digital Reading Limiting Reactants Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactants Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactants Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactants Answer Key

- Personalized Recommendations
 - Limiting Reactants Answer Key User Reviews and Ratings
 - Limiting Reactants Answer Key and Bestseller Lists
5. Accessing Limiting Reactants Answer Key Free and Paid eBooks
 - Limiting Reactants Answer Key Public Domain eBooks
 - Limiting Reactants Answer Key eBook Subscription Services
 - Limiting Reactants Answer Key Budget-Friendly Options
 6. Navigating Limiting Reactants Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactants Answer Key Compatibility with Devices
 - Limiting Reactants Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactants Answer Key
 - Highlighting and Note-Taking Limiting Reactants Answer Key
 - Interactive Elements Limiting Reactants Answer Key
 8. Staying Engaged with Limiting Reactants Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactants Answer Key
 9. Balancing eBooks and Physical Books Limiting Reactants Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactants Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Limiting Reactants Answer Key
 - Setting Reading Goals Limiting Reactants Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Limiting Reactants Answer Key

- Fact-Checking eBook Content of Limiting Reactants Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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