

### Limiting & Excess Reactants

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Cheng, C., Hansen, T., McClelland, W., Salazar, M., Shapoval, P., & Warner, J. (2003). *Estuaries of Canada: A National System of Estuaries and Watersheds*.

100

21. **Redox reactions:**  $\text{Fe}^{2+} + \text{H}_2\text{O}_2 + 2\text{H}^+ \rightarrow \text{Fe}^{3+} + 2\text{H}_2\text{O}$   
Identify the oxidizing and reducing agents.  
Solution:  $\text{Fe}^{2+}$  is oxidized to  $\text{Fe}^{3+}$  and  $\text{H}_2\text{O}_2$  is reduced to  $\text{H}_2\text{O}$ .  
Oxidizing agent:  $\text{H}_2\text{O}_2$ , Reducing agent:  $\text{Fe}^{2+}$
  22. **Electrochemistry:** Calculate the standard cell potential for the following cell:  
 $\text{Zn} | \text{Zn}^{2+} (1\text{M}) || \text{Cu}^{2+} (1\text{M}) | \text{Cu}$   
Solution:  $E^\circ_{\text{cell}} = E^\circ_{\text{cathode}} - E^\circ_{\text{anode}} = 0.34\text{V} - (-0.76\text{V}) = 1.10\text{V}$
  23. **Chemical Equilibrium:** For the reaction  $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$ ,  $K_c = 4.0$  at 400°C. Calculate  $K_p$  at the same temperature.  
Solution:  $K_p = K_c (RT)^{\Delta n} = 4.0 (0.0821 \times 400)^{-2} = 0.016$
  24. **Thermodynamics:** Calculate  $\Delta G^\circ$  for the reaction  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$  at 298 K.  
Solution:  $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ = -393.5\text{ kJ/mol} - (298\text{ K})(-0.094\text{ kJ/mol}\cdot\text{K}) = -381.5\text{ kJ/mol}$
  25. **Acid-Base Chemistry:** Calculate the pH of a 0.1 M  $\text{CH}_3\text{COOH}$  solution.  $K_a = 1.8 \times 10^{-5}$ .  
Solution:  $\text{pH} = 3.38$
  26. **Organic Chemistry:** Propose a mechanism for the reaction of  $\text{CH}_3\text{Br}$  with  $\text{OH}^-$ .  
Solution:  $\text{CH}_3\text{Br} + \text{OH}^- \rightarrow \text{CH}_3\text{OH} + \text{Br}^-$
  27. **Physical Chemistry:** Calculate the boiling point of water at 1 atm pressure.  
Solution:  $100^\circ\text{C}$
  28. **Environmental Chemistry:** Discuss the greenhouse effect.  
Solution: The greenhouse effect is caused by the trapping of heat by greenhouse gases in the atmosphere.
  29. **Materials Science:** Explain the difference between crystalline and amorphous solids.  
Solution: Crystalline solids have a long-range order, while amorphous solids lack long-range order.
  30. **Biotechnology:** Discuss the applications of recombinant DNA technology.  
Solution: Applications include gene cloning, protein production, and gene therapy.

# Limiting And Excess Reactants Extension Questions Pogil

**Lifelique**

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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