Limiting Reagents and Percentage Yield Worksheet

1. Consider the reaction

Jedine de

$$I_2O_5(g) + 5 CO(g) ----> 5 CO_2(g) + I_2(g)$$

a) 80.0 grams of iodine(V) oxide, I₂O₅, reacts with 28.0 grams of carbon monoxide, CO.

Determine the mass of iodine I2, which could be produced?

- b) If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.
 - i) what mass of iodine was produced?
 - ii) what percentage yield of iodine was produced.
- 2. Zinc and sulphur react to form zinc sulphide according to the equation.

$$Zn + S \longrightarrow ZnS$$

If 25.0 g of zinc and 30.0 g of sulphur are mixed,

- a) Which chemical is the limiting reactant?
- b) How many grams of ZnS will be formed?
- c) How many grams of the excess reactant will remain after the reaction is over?
- 3. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen? What mass is in excess? What mass of MgO is formed?
- 4. How many grams of Al₂S₃ are formed when 5.00 grams of Al is heated with 10.0 grams S?
- 5. When MoO3 and Zn are heated together they react

$$3 \text{ Zn(s)} + 2 \text{ MoO}_3(s) -----> \text{Mo}_2\text{O}_3(s) + 3 \text{ ZnO(s)}$$

What mass of ZnO is formed when 20.0 grams of MoO₃ is reacted with 10.0 grams of Zn?

- 6. Silver nitrate, AgNO₃, reacts with ferric chloride, FeCl₃, to give silver chloride, AgCl, and ferric nitrate, Fe(NO₃)₃. In a particular experiment, it was planned to mix a solution containing 25.0 g of AgNO₃ with another solution containing 45.0 grams of FeCl₃.
 - a) Write the chemical equation for the reaction.
 - b) Which reactant is the limiting reactant?
 - c) What is the maximum number of moles of AgCl that could be obtained from this mixture?

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Marian DeWane, Heather Hattori

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