

Preparation and Standardization of a NaOH Solution and Determination of the Molar Mass of an Unknown Acid Lab Report

This experiment involves two separate acid-base titration procedures. The purpose of the first one, Preparation and Standardization of a NaOH Solution is to perform an acid-base titration, in order to determine the equivalence point of acid and base. After the acid has been neutralized by the base, the results can be used to determine the concentration of sodium hydroxide solution by using stoichiometry. For the second experiment of Determination of the Molar Mass of an Unknown Acid it was to perform an acid-base titration, in order to find the molarity and molar mass of unknown acid. The second experiment uses the concentration of NaOH solution found from the first experiment.

Titration is an experiment used often for determining the concentration of an acid and base solution through the neutralization of the standard solution of base or acid. The major point in titration is to obtain the equivalence point. The equivalence point is when the acid and base have been neutralized. it is at the equivalence point where the number of moles of NaOH and $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}_{(s)}$ are reached. At the equivalence point, the base NaOH and the acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}_{(s)}$ have canceled each other out. This can be demonstrated by the chemical formula: $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}_{(s)} + 2 \text{NaOH}_{(aq)} \rightarrow \text{Na}_2\text{C}_2\text{O}_{4(aq)} + 2 \text{H}_2\text{O}_{(aq)}$ (Titration). When the acid and base are completely neutralized and the reaction is complete, the titration has reached the endpoint. The endpoint is indicated by an indicator, which is a substance that changes color to mark the endpoint. For this experiment Phenolphthalein is the indicator, it turns light pink to show that base and acid have been neutralized and is at endpoint. Standardization is the method of knowing the exact concentration of a solution. The standard solution is the base or acid that the concentration is known. Standardization was necessary in this experiment so that the concentration of base and acid are accurate. using just the weights of base or acid doesn't give an

Lab Report For An Unknown Salt Solution

Karen Timberlake



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English in Analytical Chemistry. Communicating about Methods & Techniques. Книга для студента Надежда Зорина, Александра Соболева, 2022-10-19 The Essential Lab Manual Karen Timberlake, 2002-06-24 Drawing from the successful main Laboratory Manual the Essential Laboratory Manual includes twenty one experiments which have been revised and updated Suitable for a one or two term lab course Science Educator's Guide to Laboratory Assessment Rodney L. Doran, 2002 The book opens with an up to date discussion of assessment theory research and uses Then comes a wealth of sample assessment activities in biology chemistry physics and Earth science Keyed to the National Science Education Standards the activities include reproducible task sheets and scoring rubrics All are ideal for helping students reflect on their own learning during science lab **Chemistry in the Laboratory** James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations Chemical Investigations for Changing Times C. Alton Hassell, Paula Marshall, 2003-08 by C Alton Hassell and Paula Marshall of Baylor University Contains 44 laboratory experiments and is specifically referenced to Changing Times 10 e An Instructor s Manual 0 13 140245 X prepared by Paula Marshall is also available *General, Organic, and Biological Chemistry Study Guide and Selected Solutions* Karen C. Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems **10 in One Study Package for CBSE Chemistry Class 12 with 5 Model Papers** Disha Experts, 2017-08-29 10 in ONE CBSE Study Package Chemistry class 12 with 5 Sample Papers is another innovative initiative from Disha Publication This book provides the excellent approach to Master the subject The book has 10 key ingredients that will help you achieve success 1 Chapter Utility Score 2 All India Board 2017 Solved Paper 3 Exhaustive theory based on the syllabus of NCERT books along with the concept maps for the bird s eye view of the chapter 4 NCERT Solutions NCERT Exercise Questions 5 VSA SA LA Questions Sufficient Practice Questions divided into VSA SA LA type Numericals are also included wherever required 6 Past Years Questions Past 10 year Questions of Board Exams are also included 7 HOTS Exemplar Value based Questions High Order Thinking Skill Based Moral Value Based and Selective NCERT Exemplar Questions included 8 Chapter Test A 15 marks test of 30 min to assess your preparation in each chapter 9 Important Formulae Terms and Definitions 10 Full syllabus Sample Papers 5 papers with detailed solutions designed exactly on the latest pattern of CBSE Board **Experiments in General Chemistry** Kenneth W. Whitten, Kenneth D. Gailey, 1981 10 in One Study Package for CBSE Chemistry Class 12 with Objective Questions & 3 Sample Papers 4th Edition Disha Experts, 2020-06-20 **Basic Laboratory Principles in General Chemistry** Fitzgerald B.

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