

Experiment No. 03

Date:

Name of the experiment: Standardization of HCl Solution by NaOH Solution.

**Principle**

An acid-base titration involves the addition of a **titrant** solution to an **analyte** solution. In titrimetry, chemicals used as reference solutions known as primary standards or secondary standards.

A **primary standard** is a reagent that is extremely pure, stable, has no water of hydration, and has a high molecular weight. A **secondary standard** solution refers to a solution that has its concentration measured by titration with a primary standard solution.

The secondary standard chemicals are not absolute pure and may contain some impurities that come during their synthesis. Therefore the secondary standard chemical should standardize by using the primary pure chemicals.

In the present experiment, NaOH solution is to be used secondary standard chemical and needs to be standardized by the primary standard chemicals. Here the primary standard chemical is Oxalic acid. After being standardized by the Oxalic acid, NaOH is to be used for the standardization of HCl.

The end point of the solution will be determined by an indicator, phenolphthalein.

**Reaction:****Apparatus:**

1. Electrical balance
2. Spatula
3. Volumetric flask
4. Measuring cylinder
5. Burette
6. Pipette and pipette filler
7. Conical flask
8. Beaker
9. Funnel

**Chemicals / Reagents:**

1. 0.1 N Sodium hydroxide (NaOH) solution
2. 0.1 N Oxalic acid (dihydrate) [ $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ ]
3. 0.1 N Hydro-chloric acid (HCl) solution
4. 0.5% Phenolphthalein indicator
5. Distilled Water

**Calculation (Sample):** Full calculations required to make the reagents, for the experiment must be written down:

Gram equivalent weight of  $\text{H}_2\text{C}_2\text{O}_4 = (126.08 \div 2) \text{ gm} = 63.040 \text{ gm}$

100 mL 0.1N  $\text{H}_2\text{C}_2\text{O}_4 = (63.040 \times 100 \times 0.1) \div 1000 \text{ gm} = Y \text{ gm}$  (for 100% pure  $\text{H}_2\text{C}_2\text{O}_4$ )

**Reagents and their preparations:**

1. **0.1 N 100 mL Oxalic acid solution:** Completely dissolve 0.636 gm of Oxalic acid (99%) in a 100 mL volumetric flask with a small volume of DW. Finally adjust the volume up to 100 mL mark by adding distilled water Q.S.
2. **0.1 N 1000 mL NaOH solution:** Completely dissolve 4.12 gm of NaOH (97%) in a 1000 mL volumetric flask with a small volume of DW. Finally adjust the volume up to 1000 mL mark by adding distilled water Q.S.
3. **0.1N 100 mL HCl solution:** Take 1.14 mL of HCl (32%) and completely dissolve it in 100 mL volumetric flask in 100 mL DW.
4. **0.5% 100 mL Phenolphthalein solution:** 50 mL Ethanol+ 0.5 g Phenolphthalein+ 50 mL DW.

# Laboratory Procedure Manual Format

**D Keegan**



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