

Part A: The Density of Water

Experimental Data

	1 st Water Addition	2 nd Water Addition	3 rd Water Addition
Mass of Empty Cylinder	$W_1 = 110.909 \text{ gms}$	110.909 gms	110.909 gms
Mass of Cylinder + Water	132.467 g	154.831 gms	171.421 gms
Mass of Water only	21.558 g	43.922 gms	60.512 gms
Volume of Water	22.0 ml	44.8 ml	61.5 ml
Density of Water	21.558 g/ml	0.9813 g/ml	$\frac{60.512 \text{ gms}}{61.5 \text{ ml}} = 0.9839 \text{ g/ml}$
Average Density of Water	$= (0.9813 + 0.9859) \text{ g/ml} = 0.9819 \text{ g/ml}$		

Temperature of Water: 22°C

Data Analysis

- 1) Look up the true density of water at the temperature recorded: $22^\circ \text{C} = 0.9978$
 Obtain this value from http://physics.plymouth.edu/~johnson/courses/2012_Spring/Techniques/Exams/DensityOfWater-vs-Temp.pdf. Then use this to calculate the percent error in your average density of water. Show your work.

So percentage of Error =

$$\{(EV - TV) / TV\} \times 100$$

$$= \{(0.9819 - 0.9978) / 0.9978\} \times 100$$

Part B: The Density of Aluminum and the Thickness of Foil

Experimental Data

Table 1 – The Density of Aluminum

Mass of Empty Beaker	$(W_1) = 50.075 \text{ gms}$
Mass of Beaker and Al pellets	$(W_2) = 60.266 \text{ gms}$
Mass of Al pellets	$(W) = W_2 - W_1 = 60.266 - 50.075 = 10.191 \text{ gms}$
Initial volume of water in cylinder	$(V_1) = 32.9 \text{ ml}$
Final volume of water and Al pellets	$(V_2) = 36.2 \text{ ml}$
Volume of Al pellets	$(V) = 36.2 - 32.9 = 3.3 \text{ ml}$

Lab Report On Fluid Viscosity And Density Measurements

Lauren Gardner



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