

Properties of Acids and Bases

(Adapted from Flinn Scientific Acid - Base Test Kit I - #AP4567)

Introduction

Battery acid, stomach acid, acid rain - just a few acids in our everyday life! What does it mean when something is acidic? How do acids differ from bases? In this lab, the physical and chemical properties of a variety of acids and bases will be examined.

Concepts

- Acids vs. Bases
- Indicators
- pH

Background

Acids and bases make up two groups of substances that can be categorized by their physical and chemical properties. Let's take a look at the distinguishing properties:

1. **Taste:** Acids, such as lemons or oranges, taste sour. Bases, such as soap, taste bitter. *Note:* Taste should never be used to identify a lab chemical or unknown substance. This lab does not include a test using taste.
2. **Feel:** Acid solutions do not feel much different than water; however, they sting if they contact broken skin. Base solutions have a slippery feeling. *Note:* Lab chemicals or unknown substances should never be touched with the bare skin. This lab does not include a test using feel.
3. **Indicators:** An indicator is a chemical compound, either on a test paper or in a solution, that changes color depending on the acidity or basicity of a solution and, thus, is used to test for the presence of acids or bases. There are many different kinds of indicators. Litmus paper and phenolphthalein indicator solution are two of the most common.
 - Blue litmus paper turns red when dipped in an acid solution of pH 1-6.
 - Red litmus paper turns blue when dipped in a basic solution of pH 8-14.
 - In a neutral solution with pH 7, each paper retains its original color.
 - Phenolphthalein indicator solution turns a pink color in a basic solution but remains colorless (sometimes cloudy) in an acidic or neutral solution.
4. **Reaction with carbonates:** While most bases do not react with carbonates, acids react with carbonates to release carbon dioxide gas. For example:
$$\text{BaCO}_3 + 2\text{HCl} \rightarrow \text{BaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$$
5. **Reaction with metals:** While most bases do not react with metals, acids react with active metals to release hydrogen gas. For example:
$$\text{Fe} + 2\text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2$$

Lab Properties Of Acids And Bases Answers

R Pring



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Practical Chemistry Labs Leonard Saland,1989 Features self contained step by step activities using common materials and covering topics from food chemistry to papermaking and electrochemistry Illustrates the connection between the real world and chemistry concepts such as solutions chemistry acids and bases and more Includes teacher notes quizzes and answers to help monitor student progress **Lab Exs In Prin Med Sci** Pearson Custom Publishing,1990

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Beran,2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures *How to Prepare for College Board Achievement Tests*

Joseph A. Mascetta,1990 **Handbook of Molecular Gastronomy** Christophe Lavelle,Herve This,Alan L. Kelly,Roisin Burke,2021-06-08 Handbook of Molecular Gastronomy Scientific Foundations and Culinary Applications presents a unique overview of molecular gastronomy the scientific discipline dedicated to the study of phenomena that occur during the preparation and consumption of dishes It deals with the chemistry biology and physics of food preparation along with the physiology of food consumption As such it represents the first attempt at a comprehensive reference in molecular gastronomy along with a practical guide through selected examples to molecular cuisine and the more recent applications named note by note cuisine While several books already exist for a general audience either addressing food science in general in a light way and or dealing with modern cooking techniques and recipes no book exists so far that encompasses the

whole molecular gastronomy field providing a strong interdisciplinary background in the physics biology and chemistry of food and food preparation along with good discussions on creativity and the art of cooking Features Gives A Z coverage to the underlying science physics chemistry and biology and technology as well as all the key cooking issues ingredients tools and methods Encompasses the science and practice of molecular gastronomy in the most accessible and up to date reference available Contains a final section with unique recipes by famous chefs The book is organized in three parts The first and main part is about the scientific discipline of molecular and physical gastronomy it is organized as an encyclopedia with entries in alphabetical order gathering the contributions of more than 100 authors all leading scientists in food sciences providing a broad overview of the most recent research in molecular gastronomy The second part addresses educational applications of molecular gastronomy from primary schools to universities The third part provides some innovative recipes by chefs from various parts of the world The authors have made a particular pedagogical effort in proposing several educational levels from elementary introduction to deep scientific formalism in order to satisfy the broadest possible audience scientists and non scientists This new resource should be very useful to food scientists and chefs as well as food and culinary science students and all lay people interested in gastronomy

Pesticides in Fruits and Vegetables Susan E. Kegley, Laura Wise, 1998-08-12 Students using this laboratory manual will design their own sampling plan to answer questions about organochlorine pesticides in the food supply Concepts covered include structure solubility relationships of organic compounds extraction techniques gas chromatography and risk assessment Students will use the data they collect along with selected readings to debate the issues of pesticide use

Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction

Science in Your World: Teacher edition Jay K. Hackett, 1991

The Software Encyclopedia 2000 Bowker Editorial Staff, 2000-05

Christian Home Educators' Curriculum Manual Cathy Duffy, 1997-11

Pp/Chemistry BarCharts, Inc., 2008-06-18

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