

2.3. $\Delta H = \Delta H_{\text{f}}^{\circ}(\text{C}_6\text{H}_6) - 6\Delta H_{\text{f}}^{\circ}(\text{C})$

When we use the standard heat of formation of chemical species other than elements.

The same is used for the heat of formation of the all the products, heat of formation of the reactants.

Why do we use standard heat of formation?



1. An Elliptical molecule

Calculate $\Delta H_{\text{f}}^{\circ}$ of C_6H_6 from the following data:

Types for the product heat -

Standard $\Delta H_{\text{f}}^{\circ}$ of C_6H_6 is

$\Delta H_{\text{f}}^{\circ}$ of H_2O is -285.8 kJ/mol

2. Calculate the heat of formation of the following:



The chemistry from the cell they the products not to calculate up products but the.

Para 1: $\Delta H_{\text{f}}^{\circ}(\text{C}_6\text{H}_6) = 48$

Para 2: $\Delta H_{\text{f}}^{\circ}(\text{H}_2\text{O}) = -285.8$

Para 3: $\Delta H_{\text{f}}^{\circ}(\text{CO}_2) = -393.5$

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